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## Student Exploration: Calorimetry Lab

**Vocabulary:** calorie, calorimeter, joule, specific heat capacity

### Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



- On the SIMULATION pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the GRAPH tab, and click **Play** (▶).
  - What was the **Final temperature** of the copper and the water?  
34.96
  - How much did the temperature of the copper change?  
55.04
  - How much did the temperature of the water change?  
4.96
- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Water has a greater specific heat because water resists temperature better than copper.

# Answer Sheet For Calorimetry Lab

**DJ Losen**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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