

Question 1

1(a)(iii) Difference: With powdered CaCO_3 the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid.
Similarity: The total volume of CO_2 produced is the same in both cases since the same mass (and therefore number of moles) of CaCO_3 was used in both experiments.



1(a)(v) Moles of CaCO_3 used = $(1.8/100)$ mol. = 0.018 mol.
From the balanced equation for the reaction, 1 mole CaCO_3 produces 1 mol CO_2 .
Hence volume of gas produced at R.T.P. = (0.018×24000) cm^3 = 432 cm^3 .

1(a)(vi) There were impurities in the CaCO_3 used.

1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.

1(b)(i) Fe^{3+} present.

**1(b)(ii) Magnesium ribbon decreases in size/dissolves.
Solution becomes pale green.**



1(b)(iv) SO_4^{2-} , SO_3^{2-} and CO_3^{2-} absent.

1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

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inorganic chemistry and organometallic chemistry He has Published four research papers in reputed national and international journals He has more than four years of teaching experience in various engineering colleges

Engineering Chemistry-I (For 1st Semester of Anna University) Arun Luiz T., Engineering Chemistry I

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