

Name: _____

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55 multiple-choice questions (3 pts each) + 6 fill-in questions (40 pts) = Total (205 pts).

1) How many **unpaired** electrons are present in the electron configuration of the **O** atom?

- A) none B) one
- C) two**
- D) three E) four

2) Which of the following contain(s) **polar covalent** bonds?

- A)
- H_2
- B)
- NaI
- C)
- Cl_2
- D) H_2O**
- E) both A and C

3) What is the formal charge and hybridization for the highlighted **Carbon**?

- A) 0,
- sp^2
- B) +1,
- sp^2
- C) -1, sp**
- D) +1,
- sp
- E) 0,
- sp

4) What of the following best describes the hybridization of C in **carbocation** CH_3^+ ?

- A)
- sp^3
- B) sp^2**
- C)
- sp
- D)
- s

5) what orbital is the nitrogen lone pair residing in the following molecule?

- A)
- sp^3
- B) sp^2**
- C)
- sp
- D)
- p
- E)
- s

6) Identify the **most acidic** carboxylic acid.**Answer is E**7) The **conjugate acid** of H_2O is _____

- A)
- HO^-
- B)
- H_3O
- C) H_3O^+**
- D)
- H_2O
- E)
- H_2O^+

8) How many potential stereoisomers exist for the following molecule.

- A) 1 B) 2
- C) 3**
- D) 4 E) 5



9) what is the name of the following compound?

- A) *trans*-4-ethylcyclohexan-1-ol B) *cis*-4-ethylhexan-1-ol
 C) *trans*-4-methylcyclohexan-1-ol **D) *cis*-4-ethylcyclohexan-1-ol**
 E) *trans*-4-ethylcyclopentan-1-ol



10) which of the following step is part of the propagation in the methane chlorination?

- A) $Cl-Cl \longrightarrow 2 Cl \cdot$ B) $H_3C \cdot + Cl-Cl \longrightarrow H_3C-Cl + Cl \cdot$
 C) $H_3C-H + \cdot Cl \longrightarrow H_3C \cdot + H-Cl$ D) B and C

Answer is D

11) Based on the Hammond postulate, which of the following transition state best explains the following reaction step?

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M Mark



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