

6. Given the following half-reaction reduction of a galvanic cell:



- a. Identify the anode and cathode half-reactions:



Note: This problem is an iron electrode with iron ions only in solution.

- b. Write the overall balanced cell reaction:



6. Write the cell notation for a galvanic cell that uses the reactions below:



7. A galvanic cell is constructed based on the following reaction:



- a. Calculate the standard free energy change ( $\Delta G^{\circ}$ ) for this reaction at 298 K.



$$= -2(96485 \text{ C/mol})E^{\circ}_{\text{cell}}$$

$$n = 2 \text{ mol e}^{-}$$



$$n = 2 \text{ mol e}^{-} \quad E^{\circ}_{\text{cell}} = 0.77 \text{ V} - 0.44 \text{ V} = 0.33 \text{ V}$$

$$n = 2 \text{ mol e}^{-}$$

$$n = 2 \text{ mol e}^{-}$$

- b. Is this reaction spontaneous?

Yes, this reaction is spontaneous at standard state conditions, as indicated by the negative sign of  $\Delta G^{\circ}$ .

# Answer Keys To Problem Of The Month

**Lei Huang**



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