

Mixed Gas Laws Worksheet - Solutions

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11 \text{ moles of gas}$$
- 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25°C , what will the pressure of the resulting mixture of gases be? $25^\circ\text{C} = 298 \text{ K}$

$$\text{O}_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$\text{N}_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{Tot}} = P_{\text{O}_2} + P_{\text{N}_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$
- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C . By noon the temperature has risen to 45.0°C . What is the new volume of the balloon?
 $T_1 = 20.0^\circ\text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ\text{C} = 318 \text{ K}, V_2 = ?$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$
- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(315 \text{ K})} = 260 \text{ moles of gas}$$
- 5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in $^\circ\text{C}$ of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})} = 296 \text{ K} = 23^\circ\text{C}$$
- 6) CaCO_3 decomposes at 1200°C to form CO_2 gas and CaO . If 25 L of CO_2 are collected at 1200°C , what will the volume of this gas be after it cools to 25°C ?
 $T_1 = 1200^\circ\text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ\text{C} = 298 \text{ K}, V_2 = ?$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$
- 7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at 20.0°C is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is -20.0°C ?
 $P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ\text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?,$
 $T_2 = -20.0^\circ\text{C} = 253 \text{ K}$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$
- 8) There are 135 L of gas in a container at a temperature of 260°C . If the gas was cooled until the volume decreased to 75 L, what would the temperature of the gas be?
 $T_1 = 260^\circ\text{C} = 533 \text{ K}, V_1 = 135 \text{ L}, T_2 = ?, V_2 = 75 \text{ L}$

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(75 \text{ L})(533 \text{ K})}{(135 \text{ L})} = 296 \text{ K} = 23^\circ\text{C}$$

Answer Sheet To Mixed Gas Laws

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Table of Contents Answer Sheet To Mixed Gas Laws

1. Understanding the eBook Answer Sheet To Mixed Gas Laws
 - The Rise of Digital Reading Answer Sheet To Mixed Gas Laws
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet To Mixed Gas Laws
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet To Mixed Gas Laws
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet To Mixed Gas Laws

- Personalized Recommendations
 - Answer Sheet To Mixed Gas Laws User Reviews and Ratings
 - Answer Sheet To Mixed Gas Laws and Bestseller Lists
5. Accessing Answer Sheet To Mixed Gas Laws Free and Paid eBooks
 - Answer Sheet To Mixed Gas Laws Public Domain eBooks
 - Answer Sheet To Mixed Gas Laws eBook Subscription Services
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 6. Navigating Answer Sheet To Mixed Gas Laws eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Sheet To Mixed Gas Laws Compatibility with Devices
 - Answer Sheet To Mixed Gas Laws Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet To Mixed Gas Laws
 - Highlighting and Note-Taking Answer Sheet To Mixed Gas Laws
 - Interactive Elements Answer Sheet To Mixed Gas Laws
 8. Staying Engaged with Answer Sheet To Mixed Gas Laws
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet To Mixed Gas Laws
 9. Balancing eBooks and Physical Books Answer Sheet To Mixed Gas Laws
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet To Mixed Gas Laws
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Sheet To Mixed Gas Laws
 - Setting Reading Goals Answer Sheet To Mixed Gas Laws
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Sheet To Mixed Gas Laws

- Fact-Checking eBook Content of Answer Sheet To Mixed Gas Laws
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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