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explanation the entire reaction was multiplied by 2 therefore Δh must also be halved 3 answer c explanation option c is incorrect because 566 kj mol is the Δh per mole of o2 and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes h 2 o s h 2 o l h 6 01 kj mol this is an endothermic process it requires 6 01 kj to melt one mole of ice h 2 o s the enthalpy value will change if the number of moles varies from the 1 1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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