

4.2 Solutions – Solubility Curves

Indicate the following regions (for your convenience):

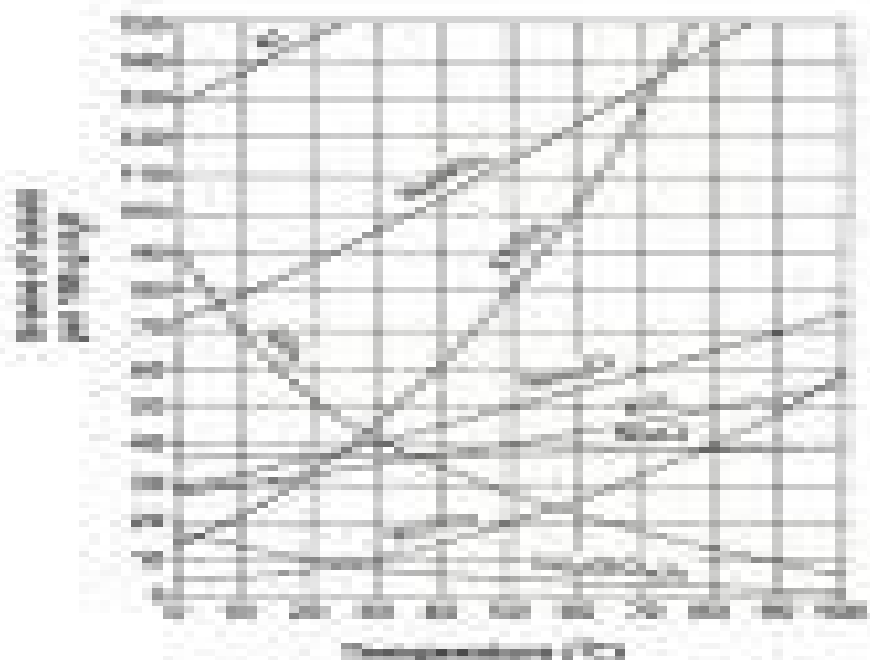
Saturated solution: Region immediately above solubility curve (no undissolved material left).

Unsaturated solution: Below solubility curve (all material is dissolved).

Supersaturated solution: Below solubility curve to right of the maximum (beyond that the solution is unstable).

Indicate the following quantities based on the solubility curve provided:

1. Dissolving 40 g of KCl in 100 g of water at 20°C .
Saturated solution
2. Mass of KCl dissolved in 100 g of water at 40°C .
60 g
3. Mass of KCl dissolved in 40 g of water at 40°C .
24 g
4. A saturated solution contains 40 g of KCl in 100 g of water at 40°C . Is this solution saturated?
Yes
5. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . Is this solution saturated?
No
6. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
7. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
8. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
9. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
10. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
11. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
12. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
13. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
14. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
15. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
16. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
17. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
18. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
19. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g
20. A saturated solution contains 40 g of KCl in 100 g of water at 20°C . What mass of KCl must be added to produce a saturated solution at 40°C ?
20 g



1 Solutions Solubility Curves Key

**Louise Petheram, Phil
Routledge, Lawrie Ryan**



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 Ranjeet Shahi, 2021-01-11

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Spotlight Science Keith Johnson, Sue Adamson, Gareth Williams, 2002 This Spiral Edition Teacher Support Pack offers comprehensive support and guidance providing the best possible learning experience for your students and saving time for everyone in the department

Clinical Aspects of Albumin S.H. Yap, C.L.H. Majoor, J.H.M. van Tongeren, 2012-12-06 Albumin is the most abundant serum protein produced by the liver In clinical practice the serum level of albumin continues to be used as an important marker of the presence progress or of the improvement of many diseases even though it is the complex end result of synthesis degradation and distribution between intra and extravascular space The clinical history of albumin began as early as in 1837 when Ansell first recognized albumen and noted that this protein is needed for transport functions for maintaining fluidity of the vascular system and for the prevention of edema However the important physiological properties

of serum proteins and their role in the regulation of the oncotic pressure were demonstrated later by the physiologist E H Starling in 1895 In 1917 the clinician A A Epstein first described the edema in patients with the nephrotic syndrome as being a result of a very low level of serum albumin Although the determination of serum albumin concentration became more popular after Howe in 1921 introduced the technique of separation of serum globulins from albumin by sodium sulfate the first preparations of human serum albumin were made available for clinical use in only 1941 by the development of plasma fractionation by Cohn and his coworkers at Harvard Medical School

Handbook of Industrial Crystallization Allan Myerson, 2002-01-08 Crystallization is an important separation and purification process used in industries ranging from bulk commodity chemicals to specialty chemicals and pharmaceuticals In recent years a number of environmental applications have also come to rely on crystallization in waste treatment and recycling processes The authors provide an introduction to the field of newcomers and a reference to those involved in the various aspects of industrial crystallization It is a complete volume covering all aspects of industrial crystallization including material related to both fundamentals and applications This new edition presents detailed material on crystallization of biomolecules precipitation impurity crystal interactions solubility and design Provides an ideal introduction for industrial crystallization newcomers Serves as a worthwhile reference to anyone involved in the field Covers all aspects of industrial crystallization in a single complete volume

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