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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	Energy	Unit
N-H	435	kJ/mol
H-O	463	kJ/mol
S-O	551	kJ/mol

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an S-O bond in kJ/mol .

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (4 \times 435) + 1(2 \times 463) \\
 & 4 \times 435 + 1(2 \times 463) \quad \underline{3150} \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{Bond broken} - \text{Bond formed} = -90 \\
 & 3150 - x = -90 \\
 & x = 3150 + 90 \\
 & \underline{3240} \text{ kJ}
 \end{aligned}$$
- Calculate the energy of an S-O bond.

$$\begin{aligned}
 & 6 \times (\text{N-H}) = 2538 \\
 & \text{N-O} = \frac{3240 - 2538}{6} \quad \underline{119} \text{ kJ/mol}
 \end{aligned}$$

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