

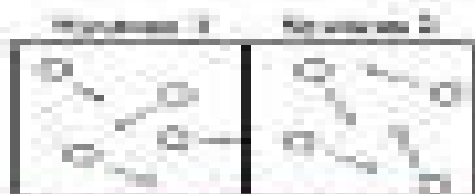
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Free-Energy Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v speed is $\frac{1}{2}mv^2$, the sum being of masses and velocities must not change, implying constant v and v temperatures are equal. The Boltzmann distribution curve shows that the average velocity is small for the distribution's maximum (low energy) and increases exponentially with temperature.

3. Sketch each graph to represent the relative distribution of the particles in the following two gas samples at the same gas:



- a. In System 1, is it a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are at a lower speed, and all the individual particles have the same mass, according to the equation $KE = \frac{1}{2}mv^2$, the velocity of all the particles shows the same mass, velocity is the same magnitude of all particles KE. In the second in System 2, the particles are moving at a faster velocity and therefore have a lower average KE.
- b. Are there different total kinetic energy for System 2? Justify your answer in terms of molecular collisions.
 All gas is transferred from System 1 to System 2. The particles are at a higher temperature in the system at a lower temperature. Therefore, the gas particles are at a lower average KE, and the total KE of the system is lower. The particles are at a lower average KE, and the total KE of the system is lower. The particles are at a lower average KE, and the total KE of the system is lower.

Answer Key Unit 6 Thermochemistry

Victor M. Corman



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