

Answer key

Unit 1 Lifestyle

Reading and Use of English: Multiple matching, pages 4–5

1

1 B 2 A 3 B 4 C 5 D 6 A 7 D 8 C 9 B
10 A

2

1 side 2 nerves 3 feet, end 4 place 5 holiday
6 lifestyle 7 own 8 shows

3

b on my own
c on our feet for hours on end
d get on each other's nerves
e putting on shows
f go on with this lifestyle

Vocabulary, pages 6–7

A Lifestyle

1 healthy 2 luxurious 3 chaotic 4 sedentary
5 alternative

B Clothes

1

1 scruffy 2 scarf 3 casual 4 tight 5 plain
6 socks 7 trainers 8 baggy 9 belt 10 helmet
11 bracelet 12 fashionable

2

1 a pleated skirt 2 a checked waistcoat
3 a flowery dress 4 striped swimming trunks
5 spotted socks

C Get

1

1 by 2 over 3 back 4 away 5 off 6 out of

2

1 touch 2 paid 3 rid 4 ready 5 worse
6 impression 7 exercise

D Word combinations

1

1 party 2 interview 3 event 4 premiere 5 life

2

1 sporting event 2 social life 3 dinner jacket
4 world premiere 5 radio interview 6 political party

Language focus, page 8

A Adverbs of frequency

1 I always set my alarm clock.
2 I am never late

3 I rarely drink it.

4 My mum cooks paella once a week.

5 we have never been to Paris.

B Be used to, get used to and used to

1 Lucy used to use a bike to get to school.

2 She's got used to getting up early in the morning.

3 My dad used to get me to clean his car every Sunday.

4 Paul didn't use to get paid much when he worked as a waiter.

5 Many young people are not used to doing hard work.

Reading and Use of English: Transformations, page 8

1 hardly ever stay/am hardly ever

2 used to getting

3 looking forward to going

4 has/s taken (me) so

5 is/s rare for Anna to

6 not like Richard/him to be

Reading and Use of English: Multiple-choice cloze, page 9

1 A 2 C 3 D 4 C 5 B 6 D 7 A 8 C

Writing: Informal letter and email, pages 10–11

1

	Formal	Informal
Complaining	7	4
Asking for information	1	10
Giving information	5	9
Apologizing	3	6
Giving advice	8	2

2

Formal	Informal
1 inform me	10 let me know
8 We strongly advise you not to	2 You really shouldn't
3 for the delay in responding to you	6 it's taken me so long to get back to you
7 Moreover	4 And
5 I have a wide range of experience in working	9 I've done loads of jobs in hotels

Informal letter

2

Paragraph 1 a Paragraph 2 c Paragraph 3 b

Answer Keys For Flyers 2 Teachers

Debbie Routh

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Let's Talk Level 3 Teacher's Manual with Audio CD Leo Jones, 2008-02-25 Let's Talk Second edition is a speaking and listening course that takes students from a high beginning to a high intermediate level The Let's Talk 3 Second Edition Teacher's Manual has been enhanced and expanded to offer increased support and flexibility Included are detailed teaching notes clear learning objectives for every activity teaching tips expansion activities and writing options Provided as photocopyables in the back of the book are model conversations for discussion support talking points for additional speaking practice and a complete assessment program including quizzes and tests The Audio CD packaged with the Teacher's Manual provides all the listening sections for the assessment program

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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
what is the limiting reagent if 8.52 g are actually formed what is the

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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
4 Al + 3 O₂ $\rightarrow$ 2 Al₂O₃
b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.93 \text{ mol Al}$
 $\frac{101.964 \text{ g Al}_2\text{O}_3}{302.3 \text{ g/mol}} = 0.337 \text{ mol Al}_2\text{O}_3$
c determine the percent yield
 $\frac{26.982 \text{ g Al}_2\text{O}_3}{101.964 \text{ g Al}_2\text{O}_3} \times 100 = 26.4\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ label 3 7 3

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