

Chemistry Semester 1 Final Exam Review

1. What is a definition of Chemistry?
Study of matter
 2. Matter includes all of the following EXCEPT:
a. air **b. light** c. smoke d. water vapor
 3. What is matter?
Something that has mass and takes up space (volume)
 4. A physical change occurs when a
a. peach spoils c. copper bowl tarnishes
b. bracelet turns your wrist green **d. glue gun melts a glue stick**
 5. The particles in a solid are
a. packed closely together c. constantly in motion
b. very far apart d. able to slide past one another
- Of the following groups, 1, 2, 14, 15, 17, 18:
6. Group(s) **1, 2** contains only metals
 7. Group (s) **18** is called a Noble gases
 8. Group(s) **17** has 7 electrons on the valence shell
 9. Group(s) **15** needs to gain 3 electrons to fill a stable octet
 10. Group(s) **2** has a 2+ oxidation state
 11. Group(s) **2** has an electron dot notation of $\bullet X \bullet$
 12. Using the periodic table, an element similar to carbon would be **14**
 13. Based on their location on the periodic table, you could infer that **18/noble gases** are very unreactive
 14. The horizontal row on the periodic table is called a(n) **period**
 15. The vertical column of blocks on the periodic table is called a(n) **group/family**
 16. The elements on the zigzag line in the periodic table are **metalloids**
 17. Which is NOT a property of metal?
a. malleability c. **unreactivity**
b. ability to lose electrons d. ability to conduct heat
 18. A volume of 1 cc (cm^3) is the same as **1** mL
 19. To determine density, the quantities that must be measured are **volume** and **mass**
 20. The relationship between the mass m of a material, its volume V , and its density D is
a. $V = mD$ b. $V/m = D$ **c. $DF = m$** d. $D + V = m$
 28. The density of aluminum is 2.70 g/cm^3 . The volume of a solid piece of Aluminum is 1.50 cm^3 . Find the mass **4.05 grams**
 21. The density of pure diamond is 3.5 g/cm^3 . The mass of a diamond is 0.25 g. Find the volume= **0.71 cm^3**
 22. What statement about density is true?
a. two samples of a pure substance may have different densities if they are different sizes
b. the density of a sample depends on its location on Earth
c. a cylinder is always used to measure the volume
d. density is a physical property that remains constant for a pure substance
 23. The number of grams equal to 0.5 kg is **500 g**
 24. In a graph the **independent** variable is placed on the x axis and the **dependant** variable is placed on the y axis.
 25. The variable that you change in the lab is the **independent** variable. The variable that responds to the change is called the **dependant** variable. All other things do not change and are called **constants**
 26. At first Dalton thought the atom was solid, spherical, **indivisible**, then Rutherford discovered the **nucleus** of the atom, then the subatomic particles **protons**, **neutrons**, and **electrons** were discovered
 27. The positively charged particle is the **proton** found in the **nucleus** of the atom
 28. The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a **neutron** and is found in the **nucleus** of the atom
 29. The **electron** is negatively charged and is found in the **electron cloud/outer shell** of the atom
 30. **Bohr** is the person credited with placing electrons on levels and **Schrodinger** established the electron cloud theory.
 31. The forces that hold the particles together in the nucleus are **columbic /electrostatic** forces
 32. Isotopes are atoms of the same element that **have different number of neutrons**
 33. All isotopes of hydrogen contain **1 proton & 1 electron**
 34. What is the atomic number for Nitrogen? **7** To what subatomic particle does this correspond? **# of protons**
 35. Zinc -66 has **30** protons and **36** neutrons
 36. What is the electron configuration for sodium? **$[\text{Ne}]3s^1$** Chlorine? **$[\text{Ne}]3s^23p^5$** Neon? **$[\text{He}]2s^22p^6$**
 37. What is the dot notation for magnesium? Bromine? Argon?

Answer To 1st Semester Chemistry Exam Review

Gary S Thorpe



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web if you start with 14 82 g of ca oh 2 c a o h 2 and 16 35 g of h₂so₄ h 2 s o 4 a determine the limiting reagent b determine the number of moles of h₂o h 2 o produced c determine the number of grams of caso₄ c a s o 4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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