

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate. The procedure is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until the chloride ions are completely precipitated as silver chloride. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of chloride ions is calculated from the mass of the precipitate using the stoichiometry of the reaction.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate.

Procedure

The procedure for this experiment is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until the chloride ions are completely precipitated as silver chloride. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of chloride ions is calculated from the mass of the precipitate using the stoichiometry of the reaction.

Data Analysis

Sample #1	Sample #2
Mass of sample	Mass of sample
Mass of precipitate	Mass of precipitate
Mass of chloride ion	Mass of chloride ion

Calculation of the mass of chloride ion in the sample #1 solution:

$$\text{Mass of Cl}^- = \text{Mass of AgCl} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

Calculation of the mass of chloride ion in the sample #2 solution:

Analysis Of Chloride Experiment 7 Lab Report

Xiang Xie



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