

Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

- 3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Answer Key To Stoichiometry Test Ch 12

Neil D. Jespersen, Pamela Kerrigan



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Table of Contents Answer Key To Stoichiometry Test Ch 12

1. Understanding the eBook Answer Key To Stoichiometry Test Ch 12
 - The Rise of Digital Reading Answer Key To Stoichiometry Test Ch 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Stoichiometry Test Ch 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Stoichiometry Test Ch 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Stoichiometry Test Ch 12
 - Personalized Recommendations
 - Answer Key To Stoichiometry Test Ch 12 User Reviews and Ratings
 - Answer Key To Stoichiometry Test Ch 12 and Bestseller Lists
5. Accessing Answer Key To Stoichiometry Test Ch 12 Free and Paid eBooks
 - Answer Key To Stoichiometry Test Ch 12 Public Domain eBooks
 - Answer Key To Stoichiometry Test Ch 12 eBook Subscription Services
 - Answer Key To Stoichiometry Test Ch 12 Budget-Friendly Options
6. Navigating Answer Key To Stoichiometry Test Ch 12 eBook Formats

- ePub, PDF, MOBI, and More
- Answer Key To Stoichiometry Test Ch 12 Compatibility with Devices
- Answer Key To Stoichiometry Test Ch 12 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Stoichiometry Test Ch 12
 - Highlighting and Note-Taking Answer Key To Stoichiometry Test Ch 12
 - Interactive Elements Answer Key To Stoichiometry Test Ch 12
- 8. Staying Engaged with Answer Key To Stoichiometry Test Ch 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Stoichiometry Test Ch 12
- 9. Balancing eBooks and Physical Books Answer Key To Stoichiometry Test Ch 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Stoichiometry Test Ch 12
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key To Stoichiometry Test Ch 12
 - Setting Reading Goals Answer Key To Stoichiometry Test Ch 12
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key To Stoichiometry Test Ch 12
 - Fact-Checking eBook Content of Answer Key To Stoichiometry Test Ch 12
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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