



## Lab Data



Verify your calculation. In your set-up, do your units cancel?

|                                |        |
|--------------------------------|--------|
| Mass of magnesium (g)          | .015   |
| Moles of magnesium (mol)       | .00062 |
| Temperature of water (°C)      | 23.0   |
| Temperature of water (K)       | 296.15 |
| Vapor pressure of water (mmHg) | 21.1   |
| Barometric pressure (mmHg)     | 778.0  |

## Observations

I observe that there are bubbles that are visible.

|                                                                     |         |
|---------------------------------------------------------------------|---------|
| Volume of hydrogen gas collected (mL)                               | 15.00   |
| Volume of hydrogen gas collected (L)                                | 0.01500 |
| Pressure of hydrogen gas (mmHg)                                     | 756.9   |
| Pressure of hydrogen gas (atm)                                      | 0.9959  |
| Moles of hydrogen gas produced (mol)                                | .00690  |
| Ideal gas constant (L · atm · mol <sup>-1</sup> · K <sup>-1</sup> ) | 0.080   |

# Answers Escience Lab Ideal Gas Law

**M Mosston**



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