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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4 a a solution of molarity 1 5 m requires 1 5 mol of na to every litre of solvent

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$$m_{mix} = \frac{m_1 + m_2 + m_3 + m_4 + m_5}{V_1 + V_2 + V_3 + V_4 + V_5} \times \frac{1000}{\rho_{mix}}$$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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0.444 mol of CoCl_2 in 0.654 l of solution	98.0 g of phosphoric acid H_3PO_4 in 1.00 l of solution
0.2074 g of calcium hydroxide Ca(OH)_2 in 40.00 ml of solution	10.5 kg

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