

**Lab 10: Calculations with a Chemical Structure**

**Objectives:**

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product produced.
5. To compare the theoretical amount to the actual amount of product formed and calculate the percent yield.

**Reaction equation:**

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

**Procedure:**

**Calculations and Results:**

**1. Calculations:**

$$\text{moles of sodium chloride} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{5.00 \text{ g}}{58.44 \text{ g/mol}} = 0.0856 \text{ mol}$$

$$\text{moles of calcium chloride} = \frac{\text{mass of CaCl}_2}{\text{molar mass of CaCl}_2} = \frac{5.00 \text{ g}}{110.98 \text{ g/mol}} = 0.0451 \text{ mol}$$

$$\text{moles of sodium hydroxide} = \frac{\text{mass of NaOH}}{\text{molar mass of NaOH}} = \frac{5.00 \text{ g}}{40.00 \text{ g/mol}} = 0.125 \text{ mol}$$

$$\text{moles of product} = \frac{\text{mass of product}}{\text{molar mass of product}} = \frac{5.00 \text{ g}}{110.98 \text{ g/mol}} = 0.0451 \text{ mol}$$

**2. Chemical equation:  $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$**

**3. Reaction:**

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

$$\text{moles of NaCl} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{5.00 \text{ g}}{58.44 \text{ g/mol}} = 0.0856 \text{ mol}$$

# Answers For Lab 6b Heath Chemistry

**Ying Liu**



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