

Student Exploration: Calorimetry Lab

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the Calorimetry Lab Gizmo™ to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** ().

A. What was the **Final temperature** of the copper and the water?

34.96°C

B. How much did the temperature of the copper change?

$$\begin{aligned} T &= T_f - T_i \\ &= (34.96^\circ\text{C} - 90^\circ\text{C}) \\ &= -55.04^\circ\text{C} \end{aligned}$$

The temperature of the copper changed about 55°C.

C. How much did the temperature of the water change?

$$\begin{aligned} T &= T_f - T_i \\ &= (34.96^\circ\text{C} - 30^\circ\text{C}) \\ &= 4.96^\circ\text{C} \end{aligned}$$

The temperature of the water changed about 5°C.

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

$$\begin{aligned} c_w &= 4.184 \text{ J/g} \\ c_c &= 0.385 \text{ J/g} \end{aligned}$$

Water has a greater specific heat capacity.

Activity A:

Heat transfer

Get the Gizmo ready:

- Click **Reset** (.



Question: What factors determine how heat energy transfers between objects?

- Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.

A. How do you think increasing the water's mass would affect the final temperature?

Answers For Calorimetry Gizmo

**Stefan Mathias Sarge, Günther W. H.
Höhne, Wolfgang Hemminger**



Answers For Calorimetry Gizmo:

Calorimetry Stefan Mathias Sarge, Günther W. H. Höhne, Wolfgang Hemminger, 2014-02-25 Clearly divided into three parts this practical book begins by dealing with all fundamental aspects of calorimetry The second part looks at the equipment used and new developments The third and final section provides measurement guidelines in order to obtain the best results The result is optimized knowledge for users of this technique supplemented with practical tips and tricks

Calorimetry, 2016 **Calorimetry ...** American Meter Company, 1914 **Calorimetry** W. Hemminger, Günther Höhne, 1984 **A convection-scope & calorimeter** Alfred Rosling Bennett, 1897 **Calorimetry** Frank H. Bates, 1902

Application of Calorimetry in Life Sciences I. Lamprecht, 1977 *Animal and Human Calorimetry* J. A. McLean, G. Tobin, 2007-12-03 This volume provides a comprehensive survey of the theory practice and techniques of calorimetry as applied to the study of energy metabolism in humans and animals Calorimetry is used to estimate nutritional requirements of man and farm livestock and to evaluate different foods It is also a powerful tool used in research into fundamental nutritional and physiological life processes and in the evaluation of stresses imposed by abnormal or severe environments It is currently being applied in various branches of medical research and can be used as a diagnostic tool in hospitals for investigation of metabolic disorders The authors discuss both direct calorimetry which measures heat loss directly and indirect calorimetry where heat loss is inferred by measurement of some of the chemical byproducts of metabolism In addition guidance is provided to the instrumentation technical problems and precautions necessary to obtain accurate calorimetric measurements

The Modern Calorimeter Walter Porter White, 1928 **Calorimetry in Medicine** William Sharp McCann, 1924

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Table of Contents Answers For Calorimetry Gizmo

1. Understanding the eBook Answers For Calorimetry Gizmo
 - The Rise of Digital Reading Answers For Calorimetry Gizmo
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Calorimetry Gizmo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Calorimetry Gizmo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Calorimetry Gizmo
 - Personalized Recommendations
 - Answers For Calorimetry Gizmo User Reviews and Ratings
 - Answers For Calorimetry Gizmo and Bestseller Lists
5. Accessing Answers For Calorimetry Gizmo Free and Paid eBooks
 - Answers For Calorimetry Gizmo Public Domain eBooks
 - Answers For Calorimetry Gizmo eBook Subscription Services
 - Answers For Calorimetry Gizmo Budget-Friendly Options

6. Navigating Answers For Calorimetry Gizmo eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Calorimetry Gizmo Compatibility with Devices
 - Answers For Calorimetry Gizmo Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Calorimetry Gizmo
 - Highlighting and Note-Taking Answers For Calorimetry Gizmo
 - Interactive Elements Answers For Calorimetry Gizmo
8. Staying Engaged with Answers For Calorimetry Gizmo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Calorimetry Gizmo
9. Balancing eBooks and Physical Books Answers For Calorimetry Gizmo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Calorimetry Gizmo
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Calorimetry Gizmo
 - Setting Reading Goals Answers For Calorimetry Gizmo
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Calorimetry Gizmo
 - Fact-Checking eBook Content of Answers For Calorimetry Gizmo
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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