
Chem - Covalent Bonding: Questions & Answers

properties of non metals ✓ Ans - inability to conduct electricity - absence of ions or delocalised electrons
low boiling and melting points - generally weak intermolecular bonds, can be soft

molecules ✓ Ans - separate groups of atoms with a formula, bonded together

intermolecular bonds ✓ Ans - bonds between molecules (dispersion, dipole-dipole, hydrogen bonding) determines physical properties

intramolecular bonds ✓ Ans - bonds between atoms in a molecule (covalent, ionic eg.)

covalent bonds ✓ Ans - non metals share electrons to complete the octet rule, an atom forms as many bonds as it needs eg. carbon needs 4 electrons so forms 4 bonds

why do non metals share electrons not transfer? ✓ Ans - non metals have a relatively high electronegativity so competing to attract each others electron, shares instead

electron dot diagrams ✓ Ans - lewis diagram shows valence electrons (bonding pairs) and lone pairs

valence structure ✓ Ans - each pair of bonding electrons is shown as a line, still represents dots as lone pairs. 1 line=2 electrons

bonding pairs and lone pairs ✓ Ans - bonding pairs; shared electrons in bond
lone pairs; non bonding pairs

double covalent bond ✓ Ans - rigid, two pairs of electrons (4 electrons) are shared, represented as a double line in valence structures eg. O_2

Answers For Chem Think Covalent Bonding

Leallyn Burr Clapp



Answers For Chem Think Covalent Bonding:

Chemistry of the Covalent Bond Leallyn Burr Clapp, 1957

Answers For Chem Think Covalent Bonding Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the power of words has become more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such is the essence of the book **Answers For Chem Think Covalent Bonding**, a literary masterpiece that delves deep to the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall impact on readers.

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