

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
 If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the molar mass of the excess reactant.

4 Followed moles of product $\times$ molar mass of product.

5 The reaction below produces ammonia for use in fertilizers.
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6 How many grams of the excess reactant is produced in a left over?

7 Followed moles of product $\times$ molar mass of product.



8 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.
 $Mg + 2NaCl \rightarrow MgCl_2 + 2Na$
 If you start with 10 moles of magnesium and 2 moles of sodium, determine the limiting reactant.

9 Calculate the grams of magnesium chloride produced.

10 How many grams of the excess reactant is produced in a left over?

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ANSWER KEY INCLUDED

Adjusting To Reality Limiting Reactant Answer Key

Marcel A. Müller



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