

### Question 1

1(a)(iii) Difference: With powdered  $\text{CaCO}_3$  the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid.  
Similarity: The total volume of  $\text{CO}_2$  produced is the same in both cases since the same mass (and therefore number of moles) of  $\text{CaCO}_3$  was used in both experiments.



**1(a)(v)** Moles of  $\text{CaCO}_3$  used =  $(1.8/100)$  mol. = 0.018 mol.  
From the balanced equation for the reaction, 1 mole  $\text{CaCO}_3$  produces 1 mol  $\text{CO}_2$ .  
Hence volume of gas produced at R.T.P. =  $(0.018 \times 24000)$   $\text{cm}^3$  = 432  $\text{cm}^3$ .

1(a)(vi) There were impurities in the  $\text{CaCO}_3$  used.

1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.

1(b)(i)  $\text{Fe}^{3+}$  present.

**1(b)(ii) Magnesium ribbon decreases in size/dissolves.  
Solution becomes pale green.**



1(b)(iv)  $\text{SO}_4^{2-}$ ,  $\text{SO}_3^{2-}$  and  $\text{CO}_3^{2-}$  absent.

1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

# Answers For January 2014 Chemistry

**Arun Luiz T.**



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Sulthan Bathery Kerala in 2002 He Stood First in his College for B sc and M sc Chemistry He received his Ph D in Inorganic Chemistry from IIT Madras in the year 2010 His research interest includes phosphorus based ligands in synthetic inorganic chemistry and organometallic chemistry He has Published four research papers in reputed national and international journals He has more than four years of teaching experience in various engineering colleges **The Air Force Law Review** ,2015

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