

Experiment 10: Electrochemistry with a Standard Hydrogen Electrode (from *Basic Laboratory Experiments*, Experiment 10, 45-50)

The purpose of this experiment is to determine standard potentials for several other electrodes. In the experiment, a standard hydrogen electrode (SHE) is used as a reference electrode. The SHE is a half-cell consisting of a platinum electrode immersed in a 1.0 M H^+ solution, with H_2 gas at 1 atm pressure. The potential of the SHE is defined as 0.00 V. Other electrodes are connected to the SHE, and the cell potential is measured. The standard potential of the other electrode is then determined relative to the SHE. The standard potential of the other electrode is the potential of the cell when the concentration of all species is 1.0 M. The standard potential of the other electrode is determined by measuring the cell potential and adding the standard potential of the SHE (0.00 V).

In this experiment, you will determine the standard potentials for several electrodes and calculate the standard cell potential for several electrochemical cells.

Objectives:

1. To determine the standard potential for several electrodes.
2. To calculate the standard potential for the cell.
3. To determine the standard potential for the cell.
4. To determine the standard potential for the cell.
5. To calculate the standard potential for the cell.

Materials:

- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)
- Standard Hydrogen Electrode (SHE)

Procedure:

1. Prepare the standard hydrogen electrode (SHE).
2. Prepare the standard hydrogen electrode (SHE).
3. Prepare the standard hydrogen electrode (SHE).
4. Prepare the standard hydrogen electrode (SHE).
5. Prepare the standard hydrogen electrode (SHE).

Results:

1. Calculate the standard potential for the cell.
2. Calculate the standard potential for the cell.

Discussion:

1. Discuss the results of the experiment.
2. Discuss the results of the experiment.

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