

VEGIUM LAB

Average Atomic Mass

Objectives

- Determine the weight of each isotope of the fictitious element "VEGIUM".
- Determine the relative abundances of the isotopes of "VEGIUM".
- Calculate from experimentally-determined values, mass of "VEGIUM".

Introduction

Isotopes are atoms of the same atomic number having different masses due to different numbers of neutrons. The atomic mass of an element is a weighted average, taking into account both the mass and relative abundances of each isotope of that element as it occurs in nature. The relative abundances and masses of each atomic particles are measured in the laboratory by an instrument called a mass spectrometer. That mass spectrometer separates particles by mass and measures the mass and relative abundance of each type of particle in nature. From these data a weighted average is calculated to determine the atomic mass of the element.

Purpose

In this lab you will carry out experimentally and perform the necessary calculations to determine the mass, mass of isotopes, three isotopes of "VEGIUM" (three elements for purposes of this lab), and three types of particles. These isotopes are different weights of "VEGIUM" having different masses. Unlike real isotopes, the weights of each type of particle (neutrons, protons, and electrons) are all different slightly in mass, so you will determine the average mass of each type of particle. Then you will determine the relative abundance of each isotope, and use these data along with the isotopic masses to calculate the weighted average mass, or atomic mass, of "VEGIUM".

Materials & Equipment

- 30g of "VEGIUM"
- 100g of "VEGIUM"
- weighing boat that will fit in weighing dish

Experimental Procedure

Carry out the following steps and record your results in your data table. Round all calculations neatly to a maximum of four figures. Report all numbers to the right significant figures.

1. Weigh all three boxes. Then all three pieces. Then all three isotopes: protons, neutrons and electrons.
 2. Count all the protons. Then all the neutrons. Then all the electrons. Report your data.
 3. Divide the total mass of each isotope by the number of each isotope to get the average mass of each isotope.
- Now, in all cases except the average mass of the neutrons, the numbers in the "total" column are the "average" values. For neutrons, the numbers are the "total" values. Do not forget to divide the total mass of the neutrons by the number of neutrons to get the average mass of the neutrons.

Answers For Vegium Lab

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Addison-Wesley Small-scale Chemistry Dennis D. Staley, Edward L. Waterman, 1995

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