

Chemistry 30 - Review - Assignment 2

KEY

1. Complete and classify the following partial equations



2. Given: $2\text{NaClO}_3 \rightarrow 2\text{NaCl} + 3\text{O}_2$. Calculate the mass of oxygen gas produced if 145 g of NaClO_3 is decomposed.

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{25}\text{H}_{52}$).

834g

4. If 14.0 moles of butane (C_4H_{10}) is burned

a) calculate the number of moles of oxygen gas required.

91.0 mol O_2

b) calculate the number of moles of carbon dioxide produced

56.0 mol CO_2

c) calculate the mass of water produced.

10.0 mol H_2O

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate.

23.4g Al

b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate.

182g Zn

6. Given: $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$

a) How many grams of Cu are required to react with 3.00 g of silver nitrate?

0.561g Cu

b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu?

1.90g Ag

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Chemistry 30. Module 5. Chemistry in Balance Alberta. Distance Learning, 1994

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