

Experiment 37

Advance Study Assignment: Qualitative Analysis of Group II Cations

1. Prepare a complete flow diagram for the separation and identification of the Group II cations and put it on the report page for this experiment.
2. Write balanced net ionic equations for the following reactions:

- a. Precipitation of tin(IV) sulfide with H_2S



- b. The confirmatory test for antimony



- c. The dissolution of Bi_2S_3 in hot nitric acid



- d. The confirmatory test for copper



3. A solution that may contain Cu^{2+} , Bi^{3+} , Sn^{4+} , and/or Sb^{3+} ions is treated with thiocetamide in an acid medium. The black precipitate that forms is partly soluble in strongly alkaline solution. The precipitate that remains is soluble in 6 M HNO_3 and gives only a blue solution on treatment with excess NH_3 . The alkaline solution, when acidified, produces an orange precipitate. On the basis of this information, which ions are present, which are absent, and which are still in doubt? State your reasoning. Note: For "paper unknowns" such as this one, confirmatory tests are usually not included, and you do not need to provide any. The information provided here, without confirmatory tests, is sufficient to clearly indicate the presence or absence of some of the Group II ions studied in this experiment, while leaving others in doubt.

Present _____

Absent _____

In doubt _____

Advance Study Assignment Group 1 Cations Answers

ML Morrison

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the ML Morrison text.

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