

Question 1: Water at 20°C falls from a height of 854 m. If the whole energy is used in increasing the temperature, find out the final temperature.
Specific heat of water is 4200 J/K kg .

Given that

$$T_1 = 20^{\circ}\text{C} = 20 + 273 = 293 \text{ K}$$

$$\text{height} = h = 854 \text{ m}$$

$$\text{Specific heat of water} = 4200 \text{ J K}^{-1} \text{ kg}^{-1}$$

$$\text{Final temperature} = T_2 = ?$$

$$\text{P.E} = mgh \quad \text{--- (1)}$$

Solution:

We know that

$$\Delta Q = cm \Delta T \quad \text{--- (2)}$$

Whole energy is used in increasing the temperature (Given)

$$\Delta Q = \text{P.E}$$

From (1) and (2)

$$cm \Delta T = mgh$$

$$c \Delta T = gh$$

$$c (T_2 - T_1) = gh$$

$$T_2 - T_1 = \frac{gh}{c}$$

$$T_2 = \frac{gh}{c} + T_1$$

$$= \frac{9.8 \times 854}{4200} + 293$$

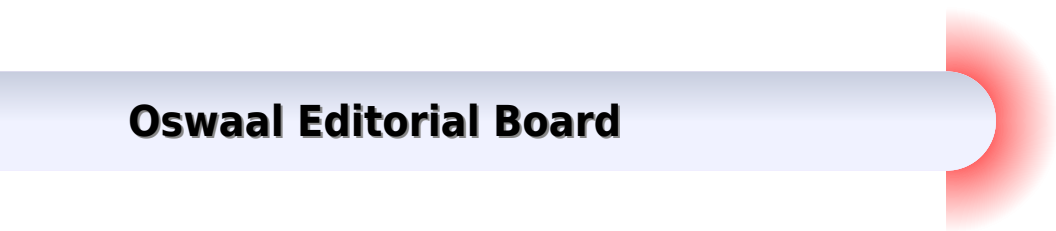
$$= \frac{8369.2}{4200} + 293 = 1.99 + 293$$

$$T_2 = 294.99 \text{ K} = 295 - 273^{\circ}\text{C}$$

$$T_2 = 22^{\circ}\text{C}$$

Answer The Following Questions About The Thermodynamics Of

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Unveiling the Magic of Words: A Review of "**Answer The Following Questions About The Thermodynamics Of**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Answer The Following Questions About The Thermodynamics Of**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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