

Mole Calculations Review Worksheet – answers on next page.

Complete the problems on this worksheet to review for the Chapter 10 test. Be sure to use dimensional analysis, and label all numbers with units when solving the problems.

- Calculate the molar mass of each compound.
 - LiOH
 - barium bromide
 - $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$
 - $\text{Ca}(\text{NO}_3)_2$
- How many molecules are in 45.0 grams CH_4 ?
- How many moles are in 18.8 grams NaOH?
- A salt shaker containing 9.58×10^{23} formula units NaCl contains how many moles?
- Which of the following has the greatest mass?
4.2 moles C or 8.34×10^{28} f.u. CaCO_3 or 12.6 g $\text{Al}(\text{NO}_3)_3$
- Find the number of representative particles in each of the following:
 - 0.28 mol Ni
 - 1.84 mol S
 - 1.32 mol F_2
 - 3.15 mol K_3PO_4
- How many liters would 1.45 moles oxygen gas occupy at STP?
- How many grams would 15.8 L of ammonia occupy at STP?
- How many total atoms are in 5.8 mol CaCl_2 ?
- Calculate the percent composition of each element in the following compounds:
 - iron(III) oxide
 - $\text{Ca}_3(\text{PO}_4)_2$
 - methane
 - H_2SO_4
- Calculate the empirical and molecular formulas of adrenaline. The compound has a molecular mass of 183 g/mol, and is 59.0% C, 7.1% H, 26.2% O, and 7.7% N by mass.
- A compound is 72.2% Mg and 27.8% N by mass. Calculate the empirical formula.
- Nicotine is 74.1% C, 8.6% H, and 17.3% N by mass. The molecular mass of nicotine is 162 g/mol. Calculate its empirical and molecular formulas.
- How many total ions are in 2.95 grams MgSO_4 ?

Answers Chemistry The Mole Review Questions

David Shubert, PhD, John Leyba, PhD



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