

CHAPTER 10: Solutions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The standard heat decomposition of a substance is determined by the following thermodynamic equation:



- a. Is energy being added to or is it being subtracted from the decomposition?
- b. What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

3. _____ $\Delta H_{\text{sub}}^\circ$ is the heat of sublimation of a solid. If the heat of sublimation of a substance is 20.4 kJ/mol , what is the heat of fusion of the substance?

4. _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If the specific heat of a substance is $0.001 \text{ J/g}^\circ\text{C}$, what is the heat of fusion of the substance?

5. The product is a molecular solid and the reactant is a ionic solid, and the reaction is exothermic. What is the heat of fusion of the product?

- a. What is the value of $\Delta H_{\text{sub}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

RC Schank



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