Acids and Bases

How do acids and bases behave in water?

Why?

Acids and bows play an important role in our lives. Numerous biological processes, industrial applications, and even environmental problems are a function of the acidity or basicity (alkalinity) of aqueous solutions. It is therefore important to understand what makes a substance behave as an acid or a base when dissolved in water. In this activity, we will explore the physical and chemical properties of acids and bases.

Model 1 - Arrhenius Acids and Bases

Conomon Name for Aqueous Solution	Chemical Formula	Found in	Tastes	Turns Liteous Paper	Conducts Electricity?	Acid or Base
Acetic acid	HCH ₂ O ₂ (aq)	Vinegar	Some	Red	Yes	Acid
Benzoie acid	HC,H,O,(aq)	Food preservative	Sour	Red	Yes	Acid
Phosphoric acid	11,PO _s (aq)	Soda pop	Sour	Red	Yes	Acres
Hydrochloric acid	HCl(aq)	Soomach acid	Sour	Red	Yes	Acid
Ciuric acid	H ₁ C ₁ H ₁ O ₂ (asp)	Cienes fraies	Sour	Red	Yes	Acid
Ascorbic sold	$H_1C_1H_2C_2(xq)$	Vitamin C	! Some	Red	Yes	Acid
Magnesium bydenside	Mg(OH) (4q)	Milk of magnesia	Bitter	Blane	Yes !	Buse
Abuninum hydroxide	Al(OHI) (aqi	Amacids	Binner	Blue	Yes	Base
Bartana byelveside	Bu(OHI) (asp) 1	Labricants	POISON	Blue	Yes	Stance
Sodium hydrocide	NaOH(aq)	Drain cleaner	LCOISON	Blute	Yes	Base

1. Refer to Model 1.

at. What is the common chemical name for Vitamin C?

4. Is Viganon C classified as an acid or a base?

Examine the properties of the Archenius acids in Model 1. Liu three properties that all Archenius acids have in common.

3. Examine the obemical formulas for the Arrhenius acids in Model 1. What feature do all the Archenius acid obemical formulas have in common?

 Examine the properties of the Airhenius bases in Model 1. List two properties that all Airhenius bases have in common.

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M Woodhall

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