

General Instructions :

Read the following instructions very carefully and strictly follow them :

- (i) This question paper comprises **39** questions. **All** questions are compulsory.
- (ii) This question paper is divided into **five** sections – **A, B, C, D** and **E**.
- (iii) **Section A** – Questions No. **1** to **20** are multiple choice questions. Each question carries **1** mark.
- (iv) **Section B** – Questions No. **21** to **26** are very short answer type questions. Each question carries **2** marks. Answer to these questions should be in the range of **30** to **50** words.
- (v) **Section C** – Questions No. **27** to **33** are short answer type questions. Each question carries **3** marks. Answer to these questions should be in the range of **50** to **80** words.
- (vi) **Section D** – Questions No. **34** to **36** are long answer type questions. Each question carries **5** marks. Answer to these questions should be in the range of **80** to **120** words.
- (vii) **Section E** – Questions No. **37** to **39** are of 3 source-based / case-based units of assessment carrying **4** marks each with sub-parts.
- (viii) There is no overall choice. However, an internal choice has been provided in some sections. Only one of the alternatives has to be attempted in such questions.

SECTION A

This section has **20** multiple choice questions (Q.No. **1** – **20**). All questions are compulsory. **20 × 1 = 20**

1. The balanced chemical equation showing reaction between quicklime and water is :

- (a) $2 \text{CaO} + \text{H}_2\text{O} \longrightarrow 2 \text{CaOH} + \text{H}_2 + \text{Heat}$
- (b) $\text{CaO} + \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + \text{H}_2 + \text{Heat}$
- (c) $\text{CaO} + \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + \text{Heat}$
- (d) $2 \text{CaO} + 3 \text{H}_2\text{O} \longrightarrow 2 \text{Ca(OH)}_3 + \text{O}_2 + \text{Heat}$

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Table of Contents 3341704 Question Paper

1. Understanding the eBook 3341704 Question Paper
 - The Rise of Digital Reading 3341704 Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying 3341704 Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 3341704 Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from 3341704 Question Paper
 - Personalized Recommendations
 - 3341704 Question Paper User Reviews and Ratings
 - 3341704 Question Paper and Bestseller Lists
5. Accessing 3341704 Question Paper Free and Paid eBooks
 - 3341704 Question Paper Public Domain eBooks
 - 3341704 Question Paper eBook Subscription Services

- 3341704 Question Paper Budget-Friendly Options
- 6. Navigating 3341704 Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - 3341704 Question Paper Compatibility with Devices
 - 3341704 Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 3341704 Question Paper
 - Highlighting and Note-Taking 3341704 Question Paper
 - Interactive Elements 3341704 Question Paper
- 8. Staying Engaged with 3341704 Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 3341704 Question Paper
- 9. Balancing eBooks and Physical Books 3341704 Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 3341704 Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 3341704 Question Paper
 - Setting Reading Goals 3341704 Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 3341704 Question Paper
 - Fact-Checking eBook Content of 3341704 Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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