

Name: _____ Per: _____
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Honors Chemistry-Smith

Average Atomic Mass Practice Problems

The answers are given in bold and italics after each problem. To receive credit, you must show work. The work must be legible and organized.

- 1) Calculate the average atomic mass of lithium to the nearest thousandth, which occurs as two isotopes that have the following atomic masses and abundances in nature: 7.30% at 6.017 amu, and 92.70% at 7.018 amu. (***6.945 amu***)

$$7.3\% \times 6.017 + 92.7\% \times 7.018 = 7.028 \text{ amu}$$

- 2) Calculate the average atomic mass of magnesium using the following data for three magnesium isotopes to 3 decimal places. (***24.447 amu***)

Isotope	mass (amu)		relative abundance
Mg-24	23.985	X	0.7870 +
Mg-25	24.986	X	0.1013 +
Mg-26	25.983	X	0.1117

$$= 24.310 \text{ amu}$$

- 3) Lithium has two naturally occurring isotopes: lithium-6 and lithium-7. If the average atomic mass of lithium is 6.941 amu, which isotope is the most abundant? How do you know?

Lithium-7

6.941 is closer to 7.

- 4) The average atomic mass of copper is 63.55 amu. If the only two isotopes of copper have masses of 62.94 amu and 64.93 amu, what are the percentages of each (2 decimal places)? (This involves algebra) (***69.35% of 62.94, 30.65% of 64.93***)

Let the percentage of 62.94 amu be X.

64.93 amu $\rightarrow$ 1-X (100% - X%)

$$62.94X + 64.93(1-X) = 63.55 \quad X = 0.6935 = 69.35\%$$

$$62.94X + 64.93 - 64.93X = 63.55 \quad 1 - 0.6935 = 30.65\%$$

$$-1.99X = -1.38$$

- 5) Naturally occurring silicon consists of three isotopes, ^{28}Si , ^{29}Si , and ^{30}Si , whose atomic masses are 27.9769, 28.9865, and 29.9838, respectively. The most abundant isotope is ^{28}Si , which accounts for 92.23% of naturally occurring silicon. Given that the observed atomic mass of silicon is 28.0855, calculate the percentages of ^{29}Si and ^{30}Si in nature. (***4.75% of ^{29}Si , 3.02% of ^{30}Si***)

$$27.9769 \times 0.9223 + 28.9865X + 29.9838(0.0777 - X) = 28.0855$$

$$= 28.085$$

$$X = 4.75\%$$

$$0.0777 - 0.0475 = 0.0302 = 3.02\%$$

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Prentice Hall Chemistry, 2000 Math for Wastewater Treatment Operators, Grades 3 And 4 John Giorgi, 2011-01-12

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