

% Composition

$$\text{Mass \%} = \frac{\text{mass of element in sample}}{\text{total mass of sample}} \times 100$$

The formulas tell us atoms and moles in a substance

H_2O = 2 atoms of H, 1 atom of O

or 2 moles of H, 1 mole of O

What is the percent composition for H_2SO_4 ?

$$\% \text{H} = \frac{2.02}{98.08} \times 100 = 2.06\% \text{ H} \quad \% \text{S} = \frac{32.06}{98.08} \times 100 = 32.69$$

$$\% \text{O} = 100 - 32.69 - 2.06 = 65.25\% \text{ O}$$

Empirical and Molecular Formulas

103 Percent Composition And Chemical Formulas Answer Key

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103 Percent Composition And Chemical Formulas Answer Key:

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