

Electron Configuration

Valance electrons	Number of electrons in outermost shell.	
shell → subshells → orbitals		
shell	K, L, M, N	$n=1, 2, \dots$
subshell	s, p, d, f, g	1s, 2s, 2p, 3s, ...
orbital	s, p _x , p _y , p _z	1s, 2s, 2p _x , 2p _y , 2p _z
Element is classified as s, p, d and f block.	depending on the orbital filled by the last electron.	

Quantitative Reasoning

Mole (Avagadro's constant)	It is a unit for telling the count, just like dozen.	1 mole = 6.02×10^{23} units
Standard Temperature and Pressure (STP)		
Avagadro's Hypothesis	Volume occupied by 1 mole of ideal gas at STP is 22.41 litres.	
Mass of neutron	1.0087 u	
Mass of proton	1.0073 u	
Mass of electron	0.00055 u	

Periodic Table Groups

Group 1	Alkali Metals	Valency +1	-ium
Group 2	Alkali Earth Metals	Valency +2	-ium
Group 17	Halogens	Valency -1	-ine
Group 18	Inert (Noble) Gases	Valency 0	-on
Diatomic nonmetals	H, Halogens, C, and O		
s-block	Group 1, 2		
p-block	Group 13 to 18		
d-block	Group 3 to 12		

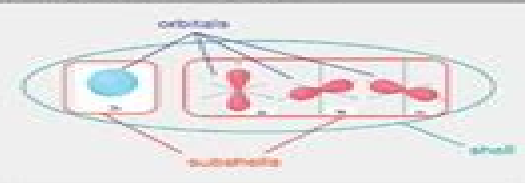
Group 1 (Alkali Metals): +1

Hydrogen (Exception: Non-metal)	H	1 (atomic number)	1x+1 u or amu
Lithium	Li	3	2x+1=7
Sodium	Na	3+8=11	2x+1=23
Potassium	K	11+8=19	2x+1=39
Rubidium	Rb	19+18=37	2x+11=85
Caesium	Cs	37+18=55	2x+23=133
Francium	Fr	55+32=87	2x+49=223

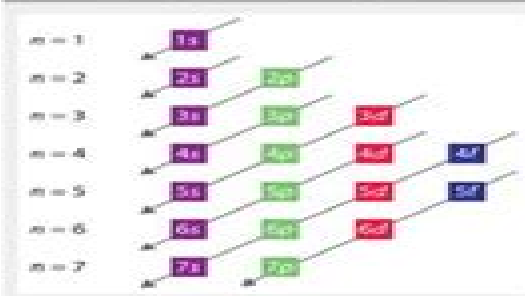
Group 2 (Alkali Earth Metals): +2

Beryllium	Be	4	
Magnesium	Mg	4+8=12	
Calcium	Ca	12+8=20	
Strontium	Sr	20+18=38	
Barium	Ba	38+18=56	
Radium	Ra	56+32=88	

Shell, Subshell and Orbital



Electron Configuration - Filling Shells



Energy levels of Orbitals



Group 17 (Halogens): -1

Fluorine	F	9	
Chlorine	Cl	9+8=17	
Bromine	Br	17+18=35	
Iodine	I	35+18=53	
Astatine (Exception: Metalloid)	At	53+32=85	
Ununseptium (Artificial Element)	Uus	85+32=117	

Group 18 (Inert or Noble Gases): 0

Helium	He	2	
Neon	Ne	2+8=10	
Argon	Ar	10+8=18	
Krypton	Kr	18+18=36	
Xeon	Xe	36+18=54	
Radon	Rn	54+32=86	
Ununoctium	Uuo	86+32=118	

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Kottai Chezhiyan



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