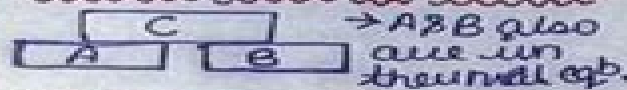


Branch of science with conversion of heat and work.

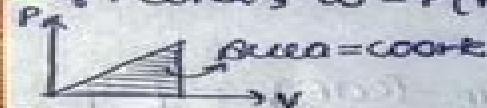
* Zeroth Law of Thermodynamics



* First Law of Thermodynamics

$$\{dQ = dw + \Delta U\}$$

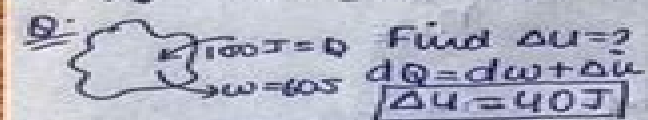
- work done $\Rightarrow dw = \int P dv$
- if P const: $w = P(V_2 - V_1)$



- $\Delta U = \frac{f}{2} n R \Delta T$
- $\{\Delta U = n C_v \Delta T\}$
- $\{\Delta U = \frac{n R \Delta T}{\gamma - 1}\}$ $\frac{f}{2} = \frac{1}{\gamma - 1}$
- $\{\Delta U = \frac{P_2 V_2 - P_1 V_1}{\gamma - 1}\}$

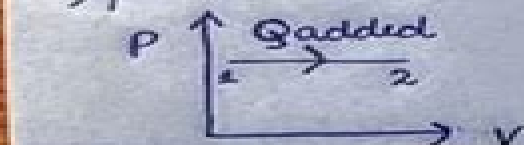
* Sign Conventions

- Heat in, work out
- $\Delta U = \frac{f}{2} n R \Delta T$
- $T \uparrow \rightarrow \Delta T \uparrow \therefore \Delta U \Rightarrow +ve$
- $T \downarrow \rightarrow \Delta T (-ve) \therefore \Delta U = -ve$



$$Q_{out} = -ve \quad Q_{in} = +ve$$

- ① Isoobaric: ($P = C$)
- $PV = nRT \rightarrow V \propto T$
- $V/T = \text{constant}$



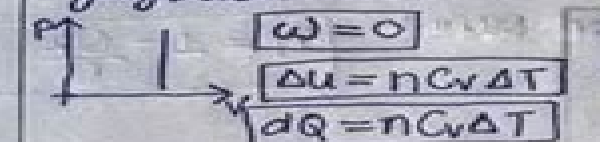
HERMODYNAMICS

$$w = \int P \cdot dv \Rightarrow w = P(V_2 - V_1)$$

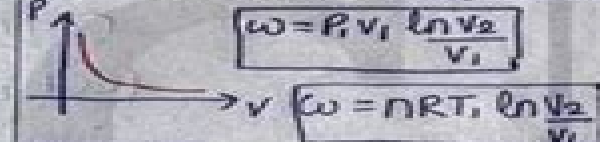
$$\Delta U = n C_v \Delta T \quad \text{for } \Delta U \text{ pure}$$

$$dQ = n C_p \Delta T$$

- (ii) Isochoric: vol const
- $P \propto T \rightarrow P/T = \text{constant}$
- Gay Lussac law.

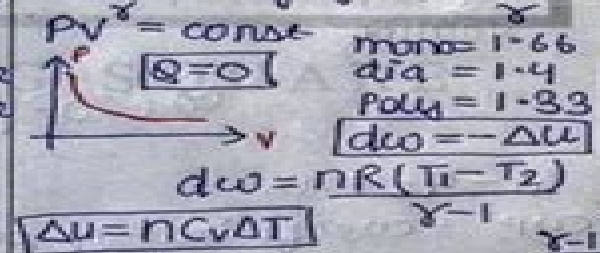


- (iii) Isothermal (Boyle's law)
- $PV = \text{constant}, xy = \text{const}$



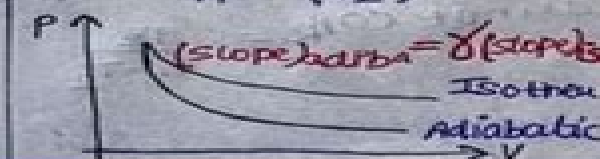
$$\Delta U = 0 \quad dQ = dw$$

- (iv) Adiabatic $Q = 0$
- Bursting of tyres



$$\frac{P_2}{P_1} = \left(\frac{V_1}{V_2}\right)^\gamma, \quad \frac{T_2}{T_1} = \left(\frac{P_2}{P_1}\right)^{\frac{\gamma-1}{\gamma}}$$

$$\frac{T_2}{T_1} = \left(\frac{V_1}{V_2}\right)^{\gamma-1}$$



(v) Polytropic:

$$PV^n = \text{const}$$

$$Q \neq 0$$

$$dQ = dw + \Delta U$$

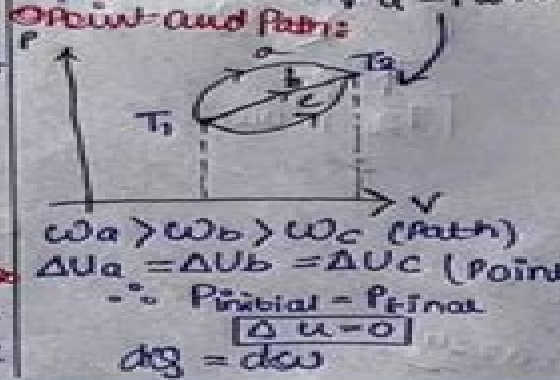
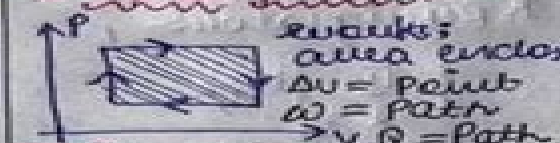
$$dw = \frac{nR(T_1 - T_2)}{\gamma - 1}$$

$$\Delta U = \frac{nRAT}{\gamma - 1} \quad \text{Polytropic Index}$$

- Or $\{\Delta U = n C_v \Delta T\}$
- $P = \text{const}$ at $n = 0$
- $PV = \text{const}$ at $n = 1$
- $PV^\gamma = \text{const}$ at $\Rightarrow n = \gamma$
- $V = \text{const}$ at $\Rightarrow n = \infty$



* Cyclic Process:



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