

Chapter 11 – Thermochemistry – Heat and Chemical Change

Chapter 11: 1 – 35, 57, 60, 61, 71

Section 11.1 – The Flow of Energy - Heat

Practice Problems

1. When 435 J of heat is added to 3.4 g of olive oil at 21°C, the temperature increases to 85°C. What is the specific heat of olive oil?

Knowns: $q = 435 \text{ J}$; $m_{\text{olive oil}} = 3.4 \text{ g}$; $\Delta T = (85^\circ\text{C} - 21^\circ\text{C})$

Use $C = q / m \cdot \Delta T$

$$C = \frac{435 \text{ J}}{3.4 \text{ g} \cdot 64^\circ\text{C}} = 2.0 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

2. A 1.55-g piece of stainless steel absorbs 141 J of heat when its temperature increases by 178°C. What is the specific heat of the stainless steel?

Knowns: $m_{\text{stainless steel}} = 1.55 \text{ g}$; $q = 141 \text{ J}$; $\Delta T = 178^\circ\text{C}$

Use $C = q / m \cdot \Delta T$

$$C = \frac{141 \text{ J}}{1.55 \text{ g} \cdot 178^\circ\text{C}} = 0.511 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

3. How much heat is required to raise the temperature of 250.0 g of mercury 52°C?

Knowns: $m_{\text{Hg}} = 250.0 \text{ g}$; $\Delta T = 52^\circ\text{C}$; $C_{\text{Hg}} = 0.14 \text{ J/g} \cdot ^\circ\text{C}$

Use $C = q / m \cdot \Delta T$; rearrange to solve for q ; $q = C \cdot m \cdot \Delta T$

$$q = \left(0.14 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}\right) (250.0 \text{ g})(52^\circ\text{C}) = 1,820 \text{ J}$$

Section Review 11.1

4. Define *energy* and explain how *energy* and *heat* are related.

Energy is the capacity for doing work or supplying heat. Heat is energy that is transferred between objects, when their temperatures are different.

5. Explain the difference between *heat capacity* and *specific heat*.

The specific heat of a substance is independent of its mass, whereas the heat capacity of an object is proportional to its mass.

6. Will the specific heat of 50 g of a substance be the same as, or greater than, the specific heat of 10 g of the same substance? Explain.

The specific heat between the two substances will be the same. Specific heat is not dependent upon the mass, whereas *heat capacity* is.

7. On a sunny day, why does the concrete deck around an outdoor swimming pool become hot, while the water stays cool?

The water stays cooler because it has a higher heat capacity than concrete. The sun's heat raises the temperature of the concrete more than that of the water.

11 Thermochemistry Heat And Chemical Change Answers

Heather Hattori, Richard H. Langley



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