# Chm 113 Final Study Guide Questions & Answers 100% Correct.

What is the [H+] of a solution with a pH of 3.78? - ANSWER14 - pOH = pH pH = 3.78 [H+] = 10^3.78 = 1.66 x 10^4

If HX represents the formula of a weak acid, what is the formula of the conjugate base of HX? - ANSWERIT HX weak acid then the formula of the conjugate base of X-

Let HX represent the formula of a weak acid. Write a balanced chemical equation for the reaction of this weak acid with the strong base NaOH? You may omit any spectator ions if you wish. - ANSWERHX + NaOH --> NaX + H2O For the equation you wrote above, what two terms are considered/assumed equal? - ANSWER[HA] = [A-]

Regarding today's experiment, what is the weak acid we are going to titrate and which strong base is used to titrate this weak acid? - ANSWERWeak Acid: Acetic Acid
Strong Base: NaOH

If a weak acid, HX, is titrated with a strong base such as NaOH, will the equivalence point in the titration occur below pH 7.0, at pH 7.0 or above pH 7.0? - ANSWERthe equivalence point in the titration will be somewhat above pH 7.0

A student is presented with a 0.65 M solution of an unknown weak acid. The student measures the pH of that solution and determines a value of 3.78 for the unknown solution. What, then, is the Ka of this weak acid? - ANSWERKa =  $\{(1.66 \times 10^{-4})^2\}/(.65 - 1.66 \times 10^{-4}) = 4.2 \times 10^{-8}$ 

What is the acid dissociation constant (Ka)? - ANSWERKa = [H30+][A-]/[HA]

What is the pH of a weak acid whose Ka is  $6.03 \times 10^{-6}$ , and the ratio of [A-]/[HA] is 3. - ANSWER6.03x10^-6 = [H3O+][9]/3 pH = 5.69

# **113 Study Guide Answers**

**Holly K. Stromberg** 

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#### **Table of Contents 113 Study Guide Answers**

- 1. Understanding the eBook 113 Study Guide Answers
  - The Rise of Digital Reading 113 Study Guide Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying 113 Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - $\circ$  Features to Look for in an 113 Study Guide Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from 113 Study Guide Answers
  - Personalized Recommendations
  - 113 Study Guide Answers User Reviews and Ratings
  - 113 Study Guide Answers and Bestseller Lists
- 5. Accessing 113 Study Guide Answers Free and Paid eBooks

- 113 Study Guide Answers Public Domain eBooks
- 113 Study Guide Answers eBook Subscription Services
- 113 Study Guide Answers Budget-Friendly Options
- 6. Navigating 113 Study Guide Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - 113 Study Guide Answers Compatibility with Devices
  - 113 Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 113 Study Guide Answers
  - Highlighting and Note-Taking 113 Study Guide Answers
  - Interactive Elements 113 Study Guide Answers
- 8. Staying Engaged with 113 Study Guide Answers
  - $\circ$  Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 113 Study Guide Answers
- 9. Balancing eBooks and Physical Books 113 Study Guide Answers
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection 113 Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 113 Study Guide Answers
  - Setting Reading Goals 113 Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 113 Study Guide Answers
  - Fact-Checking eBook Content of 113 Study Guide Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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