

11 Chemical Reactions

Big idea REACTIONS

11.1 Describing Chemical Reactions

For students using the Foundation edition, assign problems 1–17.

Essential Understanding Chemical reactions are represented by balanced chemical equations.

Lesson Summary

Introduction to Chemical Equations A chemical equation uses symbols, and sometimes words, to show the reactants and products of a chemical reaction.

- ▶ A skeleton equation uses chemical formulas to represent reactants and products, but it does not indicate the relative amounts of each.
- ▶ Anything that enters into a reaction, such as heat or a catalyst, but is not a reactant or product is shown above or below the yields arrow in the equation.

Balancing Chemical Equations Chemical equations are balanced to show that mass is conserved during chemical reactions.

- ▶ A balanced equation shows the relative amounts of reactants and products, and it contains equal numbers of each type of atom on both sides of the equation.
- ▶ Chemical equations are balanced by using coefficients in front of the chemical formulas for the reactants and the products in a skeleton equation.

BUILD Math Skills

Balancing Equations All chemical equations must be balanced because of the law of conservation of mass, which states that matter cannot be created or destroyed. So, the number of atoms that you start with at the beginning of the reaction must equal the number of atoms that you end up with.

For example, the reaction $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ follows the law of conservation of mass because you start with 2 magnesium atoms and 2 oxygen atoms and you end up with 2 magnesium atoms and 2 oxygen atoms. You can think of the $\rightarrow$ as an $=$ sign.



Turn the page to learn more about balancing equations.

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Table of Contents 11 1 Describing Chemical Reactions Workbook Answers

1. Understanding the eBook 11 1 Describing Chemical Reactions Workbook Answers
 - The Rise of Digital Reading 11 1 Describing Chemical Reactions Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 11 1 Describing Chemical Reactions Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 11 1 Describing Chemical Reactions Workbook Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 11 1 Describing Chemical Reactions Workbook Answers
 - Personalized Recommendations
 - 11 1 Describing Chemical Reactions Workbook Answers User Reviews and Ratings
 - 11 1 Describing Chemical Reactions Workbook Answers and Bestseller Lists
5. Accessing 11 1 Describing Chemical Reactions Workbook Answers Free and Paid eBooks
 - 11 1 Describing Chemical Reactions Workbook Answers Public Domain eBooks
 - 11 1 Describing Chemical Reactions Workbook Answers eBook Subscription Services
 - 11 1 Describing Chemical Reactions Workbook Answers Budget-Friendly Options
6. Navigating 11 1 Describing Chemical Reactions Workbook Answers eBook Formats

- ePub, PDF, MOBI, and More
 - 11 1 Describing Chemical Reactions Workbook Answers Compatibility with Devices
 - 11 1 Describing Chemical Reactions Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 11 1 Describing Chemical Reactions Workbook Answers
 - Highlighting and Note-Taking 11 1 Describing Chemical Reactions Workbook Answers
 - Interactive Elements 11 1 Describing Chemical Reactions Workbook Answers
8. Staying Engaged with 11 1 Describing Chemical Reactions Workbook Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 11 1 Describing Chemical Reactions Workbook Answers
9. Balancing eBooks and Physical Books 11 1 Describing Chemical Reactions Workbook Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 11 1 Describing Chemical Reactions Workbook Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 11 1 Describing Chemical Reactions Workbook Answers
- Setting Reading Goals 11 1 Describing Chemical Reactions Workbook Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 11 1 Describing Chemical Reactions Workbook Answers
- Fact-Checking eBook Content of 11 1 Describing Chemical Reactions Workbook Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} \rightarrow \text{O}_2 + 2\text{OH}^-$ 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}$ 2NO choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ h and NO_2 c h and

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