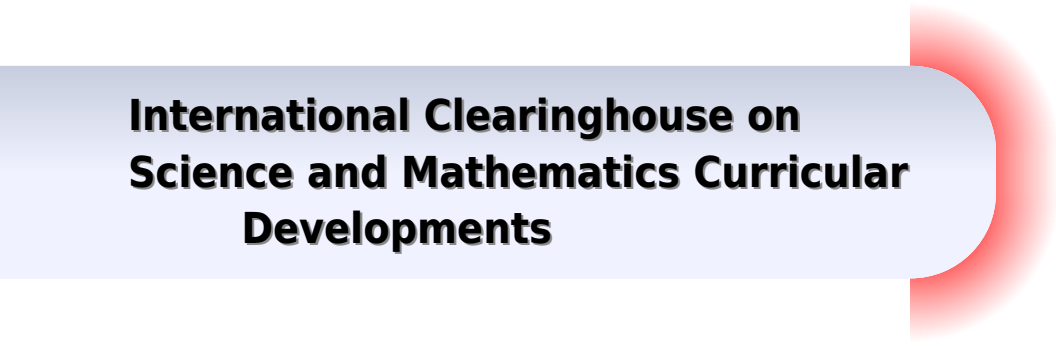

1. **Strong electrolyte:** a completely ionized or dissociated substance.
2. **Weak electrolyte (weak acid):** not completely dissociated or ionized (partial dissociation).
3. **Electrolyte:** substance that, in aqueous solution, undergoes ionization. It can therefore be said of acids or bases: solution itself is an electrolyte.
4. **Acids:** very weak or strong electrolytes (weakly or strongly ionized).
5. **Salts:** very strong or strong electrolytes (strongly or completely ionized).
6. **Law of Ostwald:** percentage ionization (or dissociation) of a weak electrolyte is inversely proportional to the square root of its concentration. It is therefore possible to determine the concentration of a weak electrolyte from its percentage ionization.
7. **Law of Ostwald:** percentage ionization (or dissociation) of a weak electrolyte is inversely proportional to the square root of its concentration. It is therefore possible to determine the concentration of a weak electrolyte from its percentage ionization.
8. **Water:** a weak electrolyte (it ionizes slightly in aqueous solution).
9. **Acid-base:** a chemical reaction in which a proton is transferred from an acid to a base.
10. **Acid-base:** a chemical reaction in which a proton is transferred from an acid to a base.
11. **Acid-base:** a chemical reaction in which a proton is transferred from an acid to a base.
12. **Acid-base:** a chemical reaction in which a proton is transferred from an acid to a base.
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