1. C₂H_{2(g)} +
$$\frac{5}{2}$$
O_{2(g)} → 2CO_{2(g)} + H₂O_(e) ΔH= -1299.6 KJ
→ 2CO_{2(g)} + H₂O_(e) → C₂H_{2(g)} + $\frac{5}{2}$ O_{2(g)}

mol_(xn)

mol_(xn)

2.
$$C_{(5)} + O_{2(9)} \rightarrow CO_{2(9)} \Delta H = -393.5 \frac{kT}{mol cxr}$$

3.
$$H_{2(g)} + \frac{1}{2}O_{2(g)} \longrightarrow H_{2(e)} \Delta H = -285.8 \text{ KJ}$$

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