

12**INTERPRETING GRAPHICS**Use with Section **6.3****Preparation of Salicylic Acid**

Student #1

mass of flask	37.820 g
flask + $C_2H_4O_2$	39.961 g
volume of $C_2H_4O_2$	5.0 mL
mass of watch glass	22.744 g
watch glass + $C_2H_4O_2$	24.489 g

Student #2

mass of flask	37.979 g
flask + $C_2H_4O_2$	40.010 g
volume of $C_2H_4O_2$	5.0 mL
mass of watch glass	21.688 g
watch glass + $C_2H_4O_2$	24.197 g

Two students prepared aspirin according to the following reaction in which acetic anhydride, $C_4H_6O_3$, reacts with salicylic acid, $C_7H_6O_3$, to form aspirin, $C_9H_8O_4$, and acetic acid, $C_2H_4O_2$.



The procedure involved heating the reaction mixture in a water bath for 15 minutes at 75°C , not to exceed 80°C . The mixture was removed from the water bath, and distilled water was added to decompose any unreacted acetic anhydride. The mixture was then placed in an ice bath for 5 minutes to facilitate the formation of aspirin crystals. The aspirin crystals were collected using filtration. The aspirin crystals were dried and then transferred to a watch glass and massed.

Because their grades were partially based on accuracy, both students used their very best lab technique. Which student got the better grade and why?

1. Determine the molar masses of the following:

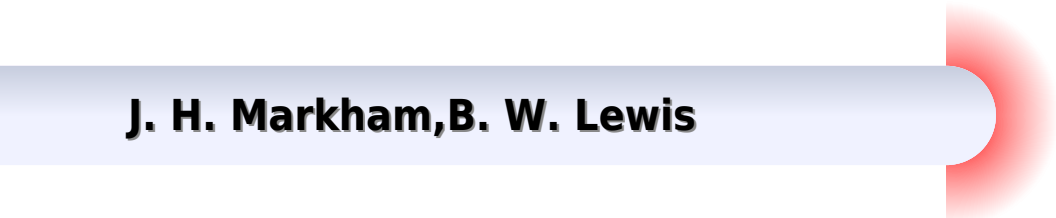
a. acetic anhydride, $C_4H_6O_3$, _____

b. salicylic acid, $C_7H_6O_3$, _____

c. aspirin, $C_9H_8O_4$, _____

12 Interpreting Graphics Preparation Of Salicylic Acid

J. H. Markham, B. W. Lewis



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The Preparation of Salicylic Acid from Phenol J. H. Markham, B. W. Lewis, 1919

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