

12 Stoichiometry



THE MOLE AND QUANTIFYING MATTER, REACTIONS

12.1 The Arithmetic of Equations

For students using the Foundation edition, assign problems 2–4, 7, 10–12.

Essential Understanding

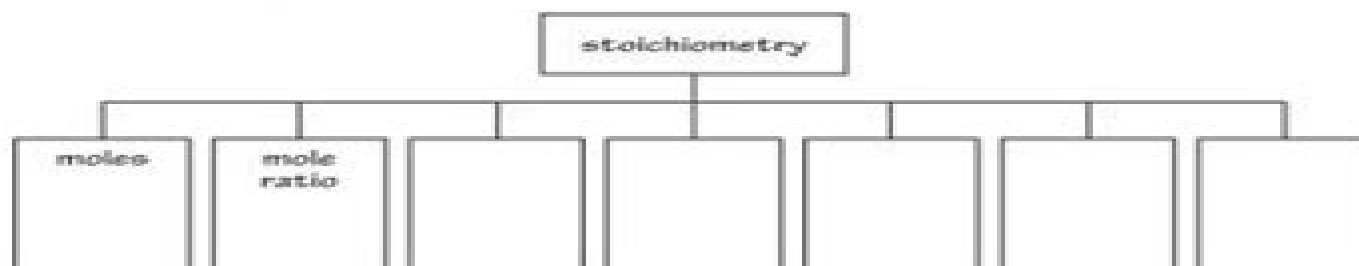
The law of conservation of mass applies to all chemical equations.



Reading Strategy

Vocabulary Word Map A vocabulary word map will help you learn vocabulary by associating the word with related words and images. Begin by writing the word *stoichiometry* in the top box.

As you read Lesson 12.1, use the word map below to help you get a better understanding of the meaning of the word *stoichiometry*. As you read, fill in the other boxes with terms, phrases, or images that are associated with the word.



Lesson Summary

Using Equations Stoichiometric calculations tell us the amounts of reactants and products under ideal conditions.

- ▶ Knowing the quantity of one substance in an equation allows you to calculate the amount of any other substance consumed or created in the reaction.

Chemical Equations The solution to every stoichiometric problem requires a balanced chemical equation.

- ▶ In a balanced chemical equation, the total number of atoms of each element in the reactants must equal the total number of atoms of that element in the products.
- ▶ In a balanced chemical equation, the total mass of the reactants must equal the total mass of the products.

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