
Reinforce and Review

Lesson Worksheets

Copy and distribute the Lesson 12.3 worksheets in CK-12 Chemistry –Intermediate Workbook. Ask students to complete the worksheets to reinforce lesson content.

Lesson Review Questions

Have students answer the Lesson Review Questions at the end of Lesson 12.3 in CK-12 Chemistry –Intermediate FlexBook® resource.

Sample Answers

Practice Problems

1.

$$1.30 \text{ mol } C_3H_8 \times \frac{5 \text{ mol } O_2}{1 \text{ mol } C_3H_8} = 6.50 \text{ mol } O_2 \text{ needed}$$

Since the moles of O_2 needed to fully react with all of the C_3H_8 is more than the amount of O_2 present (6.00 mol), O_2 is the limiting reactant.

2.

$$8.00 \text{ g } Zn \times \frac{1 \text{ mol } Zn}{65.38 \text{ g } Zn} \times \frac{2 \text{ mol } HCl}{1 \text{ mol } Zn} \times \frac{36.46 \text{ g } HCl}{1 \text{ mol } HCl} = 8.92 \text{ g } HCl \text{ needed}$$

Since there is less HCl present (8.00 g) than what is needed to react with all of the Zn, the HCl is the limiting reactant. This problem can also be solved by calculating the moles of Zn and the moles of HCl originally present. Compare the number of moles, taking note of the 1:2 ratio.

$$8.00 \text{ g } Zn \times \frac{1 \text{ mol } Zn}{65.38 \text{ g } Zn} = 0.122 \text{ mol } Zn$$
$$8.00 \text{ g } HCl \times \frac{1 \text{ mol } Zn}{36.46 \text{ g } HCl} = 0.219 \text{ mol } HCl$$

The moles of HCl is not double the moles of Zn, as required by the balanced equation. Therefore, HCl is the limiting reactant.

3.

- The balanced equation requires twice as many moles of Na as Cl_2 . Sodium (Na) is the limiting reactant.
- The 3.40 moles of Na can be converted to moles of Cl_2 that would react.

$$3.40 \text{ mol } Na \times \frac{1 \text{ mol } Cl_2}{2 \text{ mol } Na} = 1.70 \text{ mol } Cl_2 \text{ reacts}$$

The excess Cl_2 is determined by subtraction. $4.20 \text{ mol} - 1.70 \text{ mol} = 2.50 \text{ mol}$ excess Cl_2

- $3.40 \text{ mol } Na \times \frac{2 \text{ mol } NaCl}{1 \text{ mol } Na} = 3.40 \text{ mol } NaCl$

4. First, the limiting reactant must be determined.

$$35.0 \text{ g } N_2 \times \frac{1 \text{ mol } N_2}{28.02 \text{ g } N_2} \times \frac{3 \text{ mol } H_2}{1 \text{ mol } N_2} \times \frac{2.02 \text{ g } H_2}{1 \text{ mol } H_2} = 7.57 \text{ g } H_2 \text{ needed}$$

12 3 Review And Reinforce Answers

Julia Schneider



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Table of Contents **12 3 Review And Reinforce Answers**

1. Understanding the eBook **12 3 Review And Reinforce Answers**
 - The Rise of Digital Reading **12 3 Review And Reinforce Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **12 3 Review And Reinforce Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **12 3 Review And Reinforce Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **12 3 Review And Reinforce Answers**
 - Personalized Recommendations
 - **12 3 Review And Reinforce Answers** User Reviews and Ratings
 - **12 3 Review And Reinforce Answers** and Bestseller Lists

5. Accessing 12 3 Review And Reinforce Answers Free and Paid eBooks
 - 12 3 Review And Reinforce Answers Public Domain eBooks
 - 12 3 Review And Reinforce Answers eBook Subscription Services
 - 12 3 Review And Reinforce Answers Budget-Friendly Options
6. Navigating 12 3 Review And Reinforce Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 12 3 Review And Reinforce Answers Compatibility with Devices
 - 12 3 Review And Reinforce Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12 3 Review And Reinforce Answers
 - Highlighting and Note-Taking 12 3 Review And Reinforce Answers
 - Interactive Elements 12 3 Review And Reinforce Answers
8. Staying Engaged with 12 3 Review And Reinforce Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12 3 Review And Reinforce Answers
9. Balancing eBooks and Physical Books 12 3 Review And Reinforce Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12 3 Review And Reinforce Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 12 3 Review And Reinforce Answers
 - Setting Reading Goals 12 3 Review And Reinforce Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 12 3 Review And Reinforce Answers
 - Fact-Checking eBook Content of 12 3 Review And Reinforce Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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