

Name _____

5-3 Review and Reinforcement

Periodic Trends

Use the periodic table and your knowledge of periodic trends to answer the following questions.

Which atom in each pair has the larger atomic radius?

- | | |
|-----------|-------------|
| <u>K</u> | 1. Li or K |
| <u>Ca</u> | 2. Ca or Na |
| <u>Ge</u> | 3. Ge or B |
| <u>O</u> | 4. O or C |
| <u>Br</u> | 5. Cl or Br |
| <u>Ba</u> | 6. Be or Ba |
| <u>Si</u> | 7. Si or S |
| <u>Au</u> | 8. Fe or Au |

Which ion in each pair has the smaller atomic radius?

- | | |
|-----------------------------|----------------------------|
| <u>K^+</u> | 9. K^+ or Cl^- |
| <u>Be^{2+}</u> | 10. Be^{2+} or F^- |
| <u>Al^{3+}</u> | 11. Al^{3+} or P^{3-} |
| <u>K^+</u> | 12. K^+ or Ca^{2+} |
| <u>Fe^{3+}</u> | 13. Fe^{2+} or Fe^{3+} |
| <u>F^-</u> | 14. F^- or O^{2-} |

Which atom or ion in each pair has the largest ionization energy?

- | | |
|-----------|--------------|
| <u>O</u> | 15. Na or O |
| <u>Be</u> | 16. Be or Ba |
| <u>F</u> | 17. Ar or F |
| <u>Cu</u> | 18. Cu or Fe |
| <u>Ne</u> | 19. I or Ne |
| <u>V</u> | 20. K or V |
| <u>Ca</u> | 21. Ca or Fr |
| <u>Se</u> | 22. W or Se |

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JS Bruner



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