12-3 Review and Reinforcement (continued)

Solve the following problems in the space provided. Show all your work.

7. The combustion of propene proceeds in two steps:

$$CH_2CHCH_2(g) + H_2(g) \rightarrow CH_2CH_2CH_2(g)$$
 $\Delta H^2 = -124 \text{ kJ}$

$$CH_{2}CH_{2}CH_{3}(g) + 5O_{2}(g) \rightarrow 3CO_{2}(g) + 4H_{2}O(g)$$
 $\Delta H^{2} = -2220 \text{ kJ}$

Calculate the value of AH* for the combustion of 2.70 mol of propene into carbon dioxide and water.

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