

Stoichiometry Study Guide for 10th Grade

1. Molar Mass

Concept

Molar mass is the mass of one mole of a substance, measured in grams per mole (g/mol). It's a crucial concept in stoichiometry calculations.

How to Calculate

Add up the atomic masses of all atoms in a molecule or compound.

Example

Calculate the molar mass of sulphuric acid (H_2SO_4):

- 2 $\times$ H: $2 \times 1.008 \text{ g/mol} = 2.016 \text{ g/mol}$
- 1 $\times$ S: $1 \times 32.065 \text{ g/mol} = 32.065 \text{ g/mol}$
- 4 $\times$ O: $4 \times 15.999 \text{ g/mol} = 63.996 \text{ g/mol}$
- Total: $2.016 + 32.065 + 63.996 = 98.077 \text{ g/mol}$

2. Avogadro's Number

Concept

Avogadro's number is approximately 6.022×10^{23} . It represents the number of particles (atoms, molecules, or formula units) in one mole of a substance.

Example

One mole of water (H_2O) contains:

- 6.022×10^{23} water molecules
- $2 \times (6.022 \times 10^{23})$ hydrogen atoms
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