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CHEMISTRY PAPER 1
(THEORY)

Maximum Marks: 70
Time Allowed: Three hours
Candidates are allowed additional 15 minutes for only reading the paper.
They must NOT start writing during this time.

This paper is divided into four sections – A, B, C and D.
Answer all questions.
Section A consists of one question having sub-parts of one mark each.
Section B consists of ten questions of two marks each.
Section C consists of seven questions of three marks each, and
Section D consists of three questions of five marks each.
Internal choices have been provided in one question each in Section B, Section C and Section D.
All working, including rough work, should be done on the same sheet as, and adjacent to the rest of the answer.
The intended marks for questions or parts of questions are given in brackets [ ].
Balanced equations must be given wherever possible and diagrams where they are helpful.
When solving numerical problems, all essential working must be shown.
In working out problems, use the following data:
Gas constant R = 1·987 cal deg^{-1} mol^{-1} = 8·314 JK^{-1} mol^{-1}
= 0·0821 dm^3 atm K^{-1}mol^{-1}
1 l atm = 1 dm^3 atm = 101·3 J. 1 Faraday = 96500 coulombs.
Avogadro's number = 6·023 × 10^{23}.

SECTION A – 14 MARKS

Question 1

(A) Fill in the blanks by choosing the appropriate word(s) from those given in the brackets: **[4×1]**

[stable, low, aldehyde, unstable, 6, 4, ethane, Clemmensen's, 2, 3, carboxylic acid, high, propane, Rosenmund's]

(i) The primary alcohols are easily oxidised first into _____ and then into _____

(ii) The intermediate activated complex in a chemical reaction is highly _____ due to _____ energy.

(iii) The coordination number and oxidation state of the complex $K_4[Fe(CN)_6]$ are _____ and _____ respectively.

(iv) Propanone on reaction with zinc-amalgam in presence of conc. HCl gives _____ and the reaction is known as _____ reduction.

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