

Chemistry Notes for class 12 Chapter 3

Electrochemistry

Electrochemistry is that branch of chemistry which deals with the study of production of electricity from energy released during spontaneous chemical reactions and the use of electrical energy to bring about non-spontaneous chemical transformations.

Importance of Electrochemistry

1. Production of metals like Na, Mg, Ca and Al.
2. Electroplating.
3. Purification of metals.
4. Batteries and cells used in various instruments.

Conductors

Substances that allow electric current to pass through them are known as conductors.

Metallic Conductors or Electronic Conductors

Substances which allow the electric current to pass through them by the movement of electrons are called metallic conductors, e.g., metals.

Electrolytic Conductors or Electrolytes

Substances which allow the passage of electricity through their fused state or aqueous solution and undergo chemical decomposition are called electrolytic conductors, e.g., aqueous solution of acids, bases and salts.

Electrolytes are of two types:

1. **Strong electrolytes** The electrolytes that completely dissociate or ionise into ions are called strong electrolytes. e.g., HCl, NaOH, K_2SO_4
2. **Weak electrolytes** The electrolytes that dissociate partially ($\alpha < 1$) are called weak electrolytes, e.g., CH_3COOH , H_2CO_3 , NH_4OH , H_2S , etc.

Electrochemical Cell and Electrolytic

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Table of Contents 12th Class Chemistry Notes Cbse All Chapter

1. Understanding the eBook 12th Class Chemistry Notes Cbse All Chapter
 - The Rise of Digital Reading 12th Class Chemistry Notes Cbse All Chapter
 - Advantages of eBooks Over Traditional Books
2. Identifying 12th Class Chemistry Notes Cbse All Chapter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12th Class Chemistry Notes Cbse All Chapter
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12th Class Chemistry Notes Cbse All Chapter
 - Personalized Recommendations
 - 12th Class Chemistry Notes Cbse All Chapter User Reviews and Ratings
 - 12th Class Chemistry Notes Cbse All Chapter and Bestseller Lists
5. Accessing 12th Class Chemistry Notes Cbse All Chapter Free and Paid eBooks
 - 12th Class Chemistry Notes Cbse All Chapter Public Domain eBooks
 - 12th Class Chemistry Notes Cbse All Chapter eBook Subscription Services
 - 12th Class Chemistry Notes Cbse All Chapter Budget-Friendly Options
6. Navigating 12th Class Chemistry Notes Cbse All Chapter eBook Formats

- ePub, PDF, MOBI, and More
 - 12th Class Chemistry Notes Cbse All Chapter Compatibility with Devices
 - 12th Class Chemistry Notes Cbse All Chapter Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12th Class Chemistry Notes Cbse All Chapter
 - Highlighting and Note-Taking 12th Class Chemistry Notes Cbse All Chapter
 - Interactive Elements 12th Class Chemistry Notes Cbse All Chapter
 8. Staying Engaged with 12th Class Chemistry Notes Cbse All Chapter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12th Class Chemistry Notes Cbse All Chapter
 9. Balancing eBooks and Physical Books 12th Class Chemistry Notes Cbse All Chapter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12th Class Chemistry Notes Cbse All Chapter
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 12th Class Chemistry Notes Cbse All Chapter
 - Setting Reading Goals 12th Class Chemistry Notes Cbse All Chapter
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 12th Class Chemistry Notes Cbse All Chapter
 - Fact-Checking eBook Content of 12th Class Chemistry Notes Cbse All Chapter
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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