

CHAPTER 10 REVIEW*States of Matter***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Identify whether the descriptions below describe an ideal gas or a real gas.

- _____ **ideal gas** a. The gas will not condense because the molecules do not attract each other.
- _____ **ideal gas** b. Collisions between molecules are perfectly elastic.
- _____ **real gas** c. Gas particles passing close to one another exert an attraction on each other.

2. The formula for kinetic energy is $KE = \frac{1}{2}mv^2$.

- a. As long as temperature is constant, what happens to the kinetic energy of the colliding particles during an elastic collision?

The energy is transferred between them.

- b. If two gases have the same temperature and share the same energy but have different molecular masses, which molecules will have the greater speed?

Those with the lower molecule mass.

3. Use the kinetic-molecular theory to explain each of the following phenomena:

- a. A strong-smelling gas released from a container in the middle of a room is soon detected in all areas of that room.

Gas molecules are in constant, rapid, random motion.

- b. As a gas is heated, its rate of effusion through a small hole increases if all other factors remain constant.

As a gas is heated, each molecule's speed increases; therefore, the molecules pass through the small hole more frequently.

4. a. b, d, c, a List the following gases in order of rate of effusion, from lowest to highest. (Assume all gases are at the same temperature and pressure.)

(a) He (b) Xe (c) HCl (d) Cl₂

13 1 States Of Matter Section Review Answers

Wenbin Ji



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Table of Contents 13 1 States Of Matter Section Review Answers

1. Understanding the eBook 13 1 States Of Matter Section Review Answers
 - The Rise of Digital Reading 13 1 States Of Matter Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 13 1 States Of Matter Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 13 1 States Of Matter Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 13 1 States Of Matter Section Review Answers
 - Personalized Recommendations
 - 13 1 States Of Matter Section Review Answers User Reviews and Ratings
 - 13 1 States Of Matter Section Review Answers and Bestseller Lists
5. Accessing 13 1 States Of Matter Section Review Answers Free and Paid eBooks
 - 13 1 States Of Matter Section Review Answers Public Domain eBooks
 - 13 1 States Of Matter Section Review Answers eBook Subscription Services
 - 13 1 States Of Matter Section Review Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - 13 1 States Of Matter Section Review Answers Compatibility with Devices
 - 13 1 States Of Matter Section Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 13 1 States Of Matter Section Review Answers
 - Highlighting and Note-Taking 13 1 States Of Matter Section Review Answers
 - Interactive Elements 13 1 States Of Matter Section Review Answers
8. Staying Engaged with 13 1 States Of Matter Section Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 13 1 States Of Matter Section Review Answers
9. Balancing eBooks and Physical Books 13 1 States Of Matter Section Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 13 1 States Of Matter Section Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 13 1 States Of Matter Section Review Answers
- Setting Reading Goals 13 1 States Of Matter Section Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 13 1 States Of Matter Section Review Answers
- Fact-Checking eBook Content of 13 1 States Of Matter Section Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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