

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

13 Stoichiometry Practice Problems Page 20

Charles Henrickson



13 Stoichiometry Practice Problems Page 20:

This book delves into 13 Stoichiometry Practice Problems Page 20. 13 Stoichiometry Practice Problems Page 20 is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into 13 Stoichiometry Practice Problems Page 20, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to 13 Stoichiometry Practice Problems Page 20
- Chapter 2: Essential Elements of 13 Stoichiometry Practice Problems Page 20
- Chapter 3: 13 Stoichiometry Practice Problems Page 20 in Everyday Life
- Chapter 4: 13 Stoichiometry Practice Problems Page 20 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of 13 Stoichiometry Practice Problems Page 20. The first chapter will explore what 13 Stoichiometry Practice Problems Page 20 is, why 13 Stoichiometry Practice Problems Page 20 is vital, and how to effectively learn about 13 Stoichiometry Practice Problems Page 20.

3. In chapter 2, this book will delve into the foundational concepts of 13 Stoichiometry Practice Problems Page 20. The second chapter will elucidate the essential principles that must be understood to grasp 13 Stoichiometry Practice Problems Page 20 in its entirety.

4. In chapter 3, the author will examine the practical applications of 13 Stoichiometry Practice Problems Page 20 in daily life. This chapter will showcase real-world examples of how 13 Stoichiometry Practice Problems Page 20 can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of 13 Stoichiometry Practice Problems Page 20 in specific contexts. The fourth chapter will explore how 13 Stoichiometry Practice Problems Page 20 is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about 13 Stoichiometry Practice Problems Page 20. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of 13 Stoichiometry Practice Problems Page 20.

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