

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in two parts: first, the precipitation of the chloride ions is carried out, and then the precipitate is filtered, dried, and weighed. The results of the experiment are then used to calculate the mass of chloride ions in the original solution.

The reaction between silver nitrate and sodium chloride is as follows:

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

Procedure

The procedure for this experiment is as follows:

1. Weigh a sample of sodium chloride (NaCl) and dissolve it in water in a beaker.
2. Add a solution of silver nitrate (AgNO_3) to the beaker until the solution is cloudy.
3. Filter the solution and wash the precipitate with distilled water.
4. Dry the precipitate and weigh it.

Data Analysis

	Sample 1	Sample 2
Mass of NaCl	0.500 g	0.500 g
Mass of AgCl	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of chloride ion in the sample is calculated as follows:

$$\text{Mass of Cl}^- = \text{Mass of AgCl} - \text{Mass of Ag}^+ = 0.850 \text{ g} - 0.500 \text{ g} = 0.350 \text{ g}$$

The mass of chloride ion in the sample is calculated as follows:

Analysis Of Chloride Experiment 7 Lab Report

VM Jensen



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