The Periodic Table

н	4000																He
LF	Be										-	В	С	N	0	F	Ne
Na	Mg											All	Si	P	S	CI	Ar
K	Ca	Sc	TI	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rib	Sr	Y	Zr	Nb	Mo	Te	Ru	Bh	Pd	Ag	Cd	In	Sn	Sb	Te	1	Xe
Cs	Ва	La	Hif	Ta	W	Re	Os	ir	Pt	Au	Hg	TI	Pb	Bi	Po	At	Br
Fr	Ra	Ac	Rf	Ha	Sg	Uns	Uno	Une		-				siets s			
			Ce	Pr	Nd	Pm	Sm	Eu	Gd	ТЬ	Dy	Но	Er	Tm	Yb	Lu]
			Th	Pa	U	Nip	Pu	Am	Cm	Bk	Cf	Es	Em	Md	No	Lr	1

- The elements in periods 2 through 7 that are to the left of the zigzagged line are metals.
 List five metals.
- The elements in periods 1 through 6 that are to the right of the zigzagged line are nonmetals. List five nonmetals.
- The elements that border either side of the zigzagged line are metalloids. Metalloids have properties of both metals and nonmetals. Name two metalloids.
- The alkali metals are in group 1. They have only one electron in their outer shell. They
 are very reactive and have a low melting point. Shade the alkali metals blue.
- The nobel gases are found in group 18. Their outer energy level is filled, and they are very nonreactive, colorless gases. Shade the nobel gases red.
- Halogens have seven electrons in the outer shell. They are located in group 17 on the periodic table. Halogens react with metals to form salts. Shade the halogens green.
- 7. The transition elements are found in groups 3 through 12 and periods 4 through 7. These elements have either one or two electrons in the outer energy level. They are often used to form alloys because they are hard and have high melting points. Shade the transition elements orange.
- Alkaline earth metals are located in group 2. They have two electrons in their outer energy level. Shade the alkaline metals purple.

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